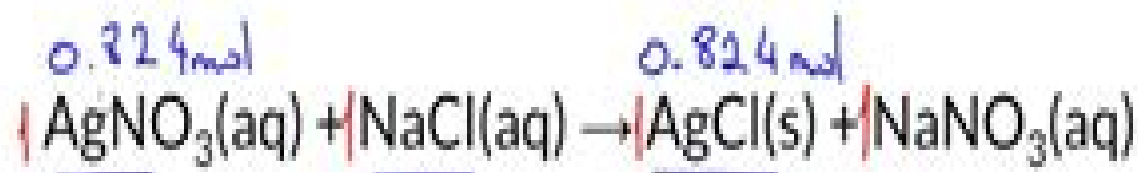


When 140.0 g of AgNO_3 was added to an aqueous solution of NaCl , 86.0 g of AgCl was collected as a white precipitate. Which salt was the limiting reactant in this reaction? How many grams of NaCl was present in the solution?



1) L.R. - ?

2) $m(\text{NaCl})$ - ?

$$n(\text{AgNO}_3) = \frac{m}{M} = \frac{140 \text{ g}}{169.9 \text{ g/mol}} = 0.824 \text{ mol}$$

$$n(\text{AgCl}) = \frac{m}{M} = \frac{86.0 \text{ g}}{143.3 \text{ g/mol}} = \underline{0.600 \text{ mol A.Y.}}$$

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Darrell Ebbing



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Table of Contents Limiting Reactant Practice Problems With Answers

1. Understanding the eBook Limiting Reactant Practice Problems With Answers
 - The Rise of Digital Reading Limiting Reactant Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Limiting Reactant Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Practice Problems With Answers
 - Personalized Recommendations
 - Limiting Reactant Practice Problems With Answers User Reviews and Ratings

- Limiting Reactant Practice Problems With Answers and Bestseller Lists
- 5. Accessing Limiting Reactant Practice Problems With Answers Free and Paid eBooks
 - Limiting Reactant Practice Problems With Answers Public Domain eBooks
 - Limiting Reactant Practice Problems With Answers eBook Subscription Services
 - Limiting Reactant Practice Problems With Answers Budget-Friendly Options
- 6. Navigating Limiting Reactant Practice Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Practice Problems With Answers Compatibility with Devices
 - Limiting Reactant Practice Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Practice Problems With Answers
 - Highlighting and Note-Taking Limiting Reactant Practice Problems With Answers
 - Interactive Elements Limiting Reactant Practice Problems With Answers
- 8. Staying Engaged with Limiting Reactant Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Practice Problems With Answers
- 9. Balancing eBooks and Physical Books Limiting Reactant Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Practice Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reactant Practice Problems With Answers
 - Setting Reading Goals Limiting Reactant Practice Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reactant Practice Problems With Answers
 - Fact-Checking eBook Content of Limiting Reactant Practice Problems With Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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