For the equilibrium, A(g)

B(g), ΔH is

40 kJ/mol. If the ratio of the activation
energies of the forward (E_f) and reverse

(E_b) reactions is $\frac{2}{3}$ then :

- (1) $E_f = 60 \text{ kJ/mol}$; $E_b = 100 \text{ kJ/mol}$
- (2) $E_f = 30 \text{ kJ/mol}$; $E_b = 70 \text{ kJ/mol}$
- (3) $E_f = 80 \text{ kJ/mol}$; $E_{fs} = 120 \text{ kJ/mol}$
- (4) $E_f = 70 \text{ kJ/mol}$; $E_b = 30 \text{ kJ/mol}$
- Determination of the molar mass of acetic acid in benzene using freezing point depression is affected by:
 - (1) dissociation
 - (2) association
 - (3) partial ionization
 - (4) complex formation
- 7. The increase of pressure on ice = water system at constant temperature will lead to:
 - (1) no effect on the equilibrium
 - (2) a decrease in the entropy of the system
 - (3) a shift of the equilibrium in the forward direction
 - (4) an increase in the Gibbs energy of the system
- 8. At 298 K, the standard reduction potentials are 1.51 V for $MnO_4^-|Mn^{2+}$, 1.36 V for $Cl_2|Cl^-$, 1.07 V for $Br_2|Br^-$, and 0.54 V for $l_2|l^-$. At pH = 3, permanganate is expected to oxidize: $\frac{RT}{F} = 0.059 \text{ V}$
 - (1) Cl⁻, Br⁻ and I⁻
 - (2) C/ and Br
 - (3) Br and I-
 - (4) 1" only

- इस साम्यावस्था के लिए, A(g) = B(g), ΔH
 -40 kJ/mol है। अगर अग्र (E_f) और पश्चगामी
 (E_b) अभिक्रियाओं के सिक्रयण कर्जाओं का अनुपात
 2
 8. तब:
 - (1) $E_f = 60 \text{ kJ/mol}$; $E_b = 100 \text{ kJ/mol}$
 - (2) $E_f = 30 \text{ kJ/mol}$; $E_b = 70 \text{ kJ/mol}$
 - (3) E_f=80 kJ/mol; E_b=120 kJ/mol
 - (4) $E_f = 70 \text{ kJ/mol}$; $E_b = 30 \text{ kJ/mol}$
- ऐसिटिक अम्ल का मोलर द्रव्यमान बेंजीन में हिमांक अवनमन से निकाले जाने पर जिससे प्रभावित होता है वह है:
 - (1) वियोजन
 - (2) संगुणन
 - (3) ऑशिक आवनीकरण
 - (4) संकुल विरचन
- स्थिर ताप पर, अधिक दाब डालने पर वर्फ

 जल समुदाय में :
 - साम्यावस्था पर कोई प्रभाव नहीं होता है।
 - (2) निकाय की ऐन्ट्रॉपी घटती है।
 - (3) साम्यावस्था अग्र दिशा में अग्रसर होती है।
 - (4) निकाय की गिष्ज कर्जा बढ़ती है।
- 298 K पर मानक अपचयन विभव, MnO₄ | Mn²⁺
 का 1.51 V, Cl₂|Cl⁻ का 1.36 V, Br₂|Br⁻ का 1.07 V और 1₂|l⁻ का 0.54 V है। pH = 3 पर MnO₄ किनका उपचयन करेगा?

$$\frac{RT}{F} = 0.059 \text{ V}$$

- (1) C7-, Br 3前t I-
- (2) CI = 新t Br =
- (3) Br = और 1 =
- (4) I THE

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Roman Wölfel

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