

Gas Variables

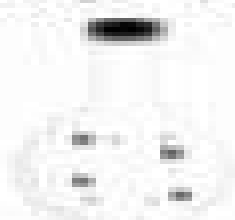
How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 1 unit

External pressure = 1 atm

Internal pressure = 1 atm

Temperature = 300 K



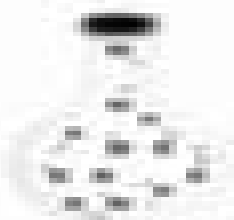
A2

Volume = 1 unit

External pressure = 1 atm

Internal pressure = 2 atm

Temperature = 300 K



A3

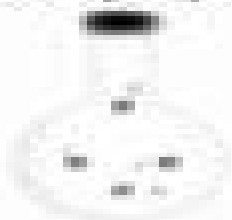
Volume = 1 unit

External pressure = 1 atm

Internal pressure = 3 atm

Temperature = 300 K

Experiment B (Heating the gas)



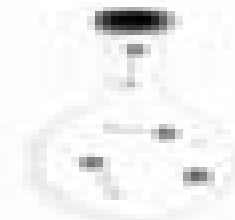
B1

Volume = 1 unit

External pressure = 1 atm

Internal pressure = 1 atm

Temperature = 300 K



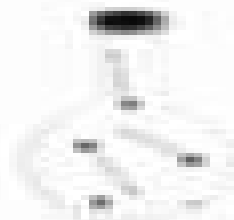
B2

Volume = 1 unit

External pressure = 1 atm

Internal pressure = 2 atm

Temperature = 400 K



B3

Volume = 1 unit

External pressure = 1 atm

Internal pressure = 3 atm

Temperature = 600 K

*Note: Volume in the model is not set in units, rather than liters, because 6 molecules of gas at the conditions given would take every third space ($1 = 33\%$) of the container. Therefore, these molecules are much larger compared to the space between them than actual gas particles.

Gas Variables Packet

Lingsheng Yao



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