

Part II: Kinetic Energy and Speed

Sketch and compare the distributions for kinetic energy and speed at two different temperatures in the table below. Record your temperatures ( $T_1$  and  $T_2$ ), set Volume as a Constant Parameter, and use roughly the same number of particles for each experiment (aim for ~100-200). Use the  $T_2$  temperature to examine a mixture of particles.

*Tips:*

$T_1 = \underline{\hspace{2cm}}$  K

*The Species Information and Energy Histograms tools will help.*

$T_2 = \underline{\hspace{2cm}}$  K

*The system is dynamic so the distributions will fluctuate.*

*Sketch the average or most common distribution that you see.*

	"Heavy" Particles Only	"Light" Particles Only	Heavy + Light Mixture
# of particles (~100-200)			
Kinetic Energy Distribution sketch for $T_1$			
Speed Distribution sketch for $T_1$			
Kinetic Energy Distribution sketch for $T_2$			
Speed Distribution sketch for $T_2$			

# Gas Properties Simulation Activity Answers

**Kathleen Armour**



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