

The Ideal and Combined Gas Laws $PV = nRT$ or $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Use your knowledge of the ideal and combined gas laws to solve the following problems. If it involves moles or grams, it must be $PV = nRT$

- 1) If four moles of a gas at a pressure of 5.4 atmospheres have a volume of 120 liters, what is the temperature?

$$(5.4)(120) = 4(.0821)T$$

$$T = 1973.2 \text{ K}$$

- 2) If I initially have a gas with a pressure of 84 kPa and a temperature of 35° C and I heat it an additional 230 degrees, what will the new pressure be? Assume the volume of the container is constant.

$$\frac{84}{(35+273)} = \frac{P_2}{(35+230+273)} \quad P_2 = 146.7 \text{ kPa}$$

- 3) My car has an internal volume of 2600 liters. If the sun heats my car from a temperature of 20° C to a temperature of 55° C, what will the pressure inside my car be? Assume the pressure was initially 760 mm Hg.

$$\frac{760 \text{ mmHg}}{(20+273)} = \frac{P_2}{(55+273)} \quad P_2 = 850 \text{ mmHg}$$

- 4) How many moles of gas are in my car in problem #3?

$$\frac{850 \text{ mmHg}}{760 \text{ mmHg}} = 1.13 \text{ atm}$$

$$2600(1.13) = n(.0821)(55+273)$$

$$n = 107.1 \text{ moles}$$

- 5) A toy balloon filled with air has an internal pressure of 1.25 atm and a volume of 2.50 L. If I take the balloon to the bottom of the ocean where the pressure is 95 atmospheres, what will the new volume of the balloon be? How many moles of gas does the balloon hold? (Assume $T = 285 \text{ K}$)

$$1.25(2.50) = 95(V_2)$$

$$V_2 = 0.032 \text{ L}$$

$$95(.032) = n(285)(.0821)$$

$$n = 0.1299 \text{ moles}$$

Gas Laws Test Answer Key

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Table of Contents Gas Laws Test Answer Key

1. Understanding the eBook Gas Laws Test Answer Key
 - The Rise of Digital Reading Gas Laws Test Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Test Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Test Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Test Answer Key
 - Personalized Recommendations
 - Gas Laws Test Answer Key User Reviews and Ratings
 - Gas Laws Test Answer Key and Bestseller Lists
5. Accessing Gas Laws Test Answer Key Free and Paid eBooks

- Gas Laws Test Answer Key Public Domain eBooks
- Gas Laws Test Answer Key eBook Subscription Services
- Gas Laws Test Answer Key Budget-Friendly Options
- 6. Navigating Gas Laws Test Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws Test Answer Key Compatibility with Devices
 - Gas Laws Test Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Test Answer Key
 - Highlighting and Note-Taking Gas Laws Test Answer Key
 - Interactive Elements Gas Laws Test Answer Key
- 8. Staying Engaged with Gas Laws Test Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Test Answer Key
- 9. Balancing eBooks and Physical Books Gas Laws Test Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Test Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Test Answer Key
 - Setting Reading Goals Gas Laws Test Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Test Answer Key
 - Fact-Checking eBook Content of Gas Laws Test Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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