

**Chemistry II-AP  
Gas Laws Practice Test  
Solution Set**

1. **B** At what temperature (on the Celsius scale) will 5.68 grams of oxygen gas exert a pressure of 335 kPa in a 1.00-liter container?  
A) 227; B) -46.1; C) -160.; D) 113; E) 718; F) 445; G) -67.2

$$PV = nRT \quad T = \frac{[335 \text{ kPa} / (101.3 \text{ kPa} / 1 \text{ atm.})][1.00 \text{ L}]}{[(5.68 \text{ g} / 32.0 \text{ g/mol})(0.0821 \text{ L-atm.} / \text{mol-K})]} = 227 \text{ K} = -46^\circ\text{C}$$

2. **A** If a balloon has a volume of 30.0 ml at STP, what will be its volume (in ml) at  $-75.0^\circ\text{C}$  and 10.0 inches Hg?  
A) 65.1; B) 32.0; C) 626; D) 125; E) 15.2; F) 6.05; G) 82.9

$$\frac{V_1 P_1}{T_1} = \frac{V_2 P_2}{T_2} \quad V_2 = \frac{(30.0 \text{ mL})(29.92 \text{ in.})}{273 \text{ K}} \times \frac{(198 \text{ K})}{10.0 \text{ in.}} = 65.10 \text{ mL}$$

3. **G** If oxygen gas has an effusion rate of 5.00 liters/min from an effusion chamber, what is the molecular weight of an unknown gas if it has an effusion rate of 3.75 liters/min?  
A) 28.4; B) 42.7; C) 18.0; D) 24.0; E) 29.9; F) 116; G) 56.9

$$\frac{\text{Rate O}_2}{\text{Rate X}} = \frac{5.00 \text{ L/min}}{3.75 \text{ L/min}} = \sqrt{\frac{\text{M.Wt. X}}{32.0 \text{ g/mol}}} \quad \text{M.Wt. X} = 56.8 \text{ g/mol}$$

4. **D** What is the density (in g/liter) of diphosphorus pentoxide gas at standard conditions?  
A) 2.16; B) 10.79; C) 1.54; D) 6.34; E) 3.25; F) 8.66; G) 5.39

$$(142.0 \text{ g/mole})(1 \text{ mole} / 22.4 \text{ L}) = 6.339 \text{ g/L}$$

5. **B** A mixture of 2.00-grams of nitrogen gas and 3.25 grams of sulfur trioxide gas are injected into a 4.00-liter container. The mixture has an equilibrium temperature of  $75.0^\circ\text{C}$ . What is the partial pressure (in atm.) of the nitrogen gas in the vessel? No reaction takes place between these two gases.  
A) 1.38; B) 0.510; C) 1.02; D) 0.872; E) 0.362; F) 0.400; G) 1.58

$$P = nRT / V = [(2.00 \text{ g} / 28.0 \text{ g/mol})(0.0821)(348 \text{ K})] / 4.00 \text{ L} = 0.51019 \text{ atm.}$$

6. **A** What happens to the mean free path of oxygen gas in the atmosphere with increasing altitude?  
A) it increases; B) it decreases; C) it remains constant

# Gas Laws Practice Test C Answer Key

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