

8. I have an unknown volume of as at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 46 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.7 \text{adm} (48) B25 K}{320 K (0.5 \text{adm})}$$

$$V_2 = L$$

 If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

I /
$$T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 L)}{323 L (3 colors)} \frac{1}{2.4 L}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

Gas Laws Practice Calculations Answer Key

Peter Atkins, Julio de Paula

Gas Laws Practice Calculations Answer Key:

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