

Gas Laws Summary

Name	Equation	Things that change	Things held constant
Combined	$P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$	Pressure, volume and temperature	Amount of gas
Boyles	$P_1 \times V_1 = P_2 \times V_2$	Pressure and Volume	Temperature Amount of gas
Charles	$V_1 \times T_2 = V_2 \times T_1$	Volume and Temperature	Pressure Amount of gas
Gay-Lussac	$P_1 \times T_2 = P_2 \times T_1$	Pressure Temperature	Volume Amount of gas

- Remember
 - Temperature MUST be in KELVIN
 - $K = ^\circ C + 273$
 - any other units MUST CANCEL
 - Can't have mL and L in same problem, can't have atm and mm Hg in same problem you MUST convert
 - $1 \text{ L} = 1000 \text{ mL}$
 - $1 \text{ atm} = 760 \text{ mm Hg} = 760 \text{ torr} = 101.3 \text{ kPa} = 101,300 \text{ Pa}$

Gas Laws Packet Answer Key

Karin Nielsen-Saines

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