

Worksheet – Gas Laws III

Gas Laws: Combined Gas Law

The following problems may be solved using the relationships described by either Charles' or Boyles' laws, together in the combined gas law. Show all the steps to the solution of the problem.

$$V_1P_1/T_1 = V_2P_2/T_2$$

1. Carbon dioxide occupies a 2.54 dm³ container at STP. What will be the volume when the pressure is 150 KPa and 26°C?

$$V_2 = \frac{[2.54 \text{ dm}^3][101.3 \text{ KPa}][26^\circ\text{C} + 273\text{K}]}{[150 \text{ KPa}][273 \text{ K}]} = 1.88 \text{ dm}^3$$

2. Oxygen occupies a fixed container of 5.5L at STP. What will happen to the pressure if the temperature rises to 300K?

$$P_2 = \frac{[5.5 \text{ L}][101.3 \text{ KPa}][300 \text{ K}]}{[5.5 \text{ L}][273 \text{ K}]} = 111.3 \text{ KPa}$$

3. Methane is compressed in a closed 15.8 dm³ container at 101.3 KPa. If the volume drops to 8.7 dm³ and the temperature begins at 25°C and then drops to 18°C , what will the pressure of the gas be?

$$P_2 = \frac{[101.3 \text{ KPa}][15.8 \text{ dm}^3][300 \text{ K}]}{[8.7 \text{ dm}^3][273 \text{ K}]} = 179.6 \text{ KPa}$$

4. A helium balloon is fully inflated at 1.2 L. When the clerk is filling the balloon, she stops to make sure it is not going to explode and checks the pressure. The pressure is 5.7 KPa and the temperature at the store is 24°C. The balloon is 0.75L full. She stops when the balloon is fully inflated and the pressure is 7.2 KPa. When the customer takes the balloon outside, it explodes. What was the temperature outdoors?

$$T = \frac{[24^\circ\text{C} + 273\text{K}][1.2 \text{ L}][7.2 \text{ KPa}]}{[0.75 \text{ L}][5.7 \text{ KPa}]} = 600 \text{ K}$$

5. Oxygen gas is added to a rebreather at the same rate as it would be in the open air. The volume is 4.8 dm³ and the temperature is 25 °C. The pressure is 1 atm. Once the tank is lowered under the water the pressure increases to 2 atm and the temperature drops to 12°C. Is the tank in danger of exploding if it's maximum volume is 5.0 dm³?

$$V_2 = \frac{[4.8 \text{ dm}^3][1 \text{ atm}][12^\circ\text{C} + 273\text{K}]}{[2 \text{ atm}][25^\circ\text{C} + 273\text{K}]} = 2.1 \text{ dm}^3, \text{ not in danger of exploding}$$

6. Ozone is formed due to electricity passing through the air and splitting some O₂ atoms which then join briefly to form O₃, ozone. Your olfactory sense can detect 1.00 mL ozone in 10 L of air at STP. You are surprised that the noreaster is accompanied by

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Table of Contents Gas Laws And Scuba Diving Chem Matters Worksheet Answers

1. Understanding the eBook Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - The Rise of Digital Reading Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Personalized Recommendations
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers User Reviews and Ratings
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers and Bestseller Lists
5. Accessing Gas Laws And Scuba Diving Chem Matters Worksheet Answers Free and Paid eBooks
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers Public Domain eBooks
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers eBook Subscription Services
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers Budget-Friendly Options
6. Navigating Gas Laws And Scuba Diving Chem Matters Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers Compatibility with Devices
 - Gas Laws And Scuba Diving Chem Matters Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Highlighting and Note-Taking Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Interactive Elements Gas Laws And Scuba Diving Chem Matters Worksheet Answers
8. Staying Engaged with Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws And Scuba Diving Chem Matters Worksheet Answers
9. Balancing eBooks and Physical Books Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws And Scuba Diving Chem Matters Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Setting Reading Goals Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Fact-Checking eBook Content of Gas Laws And Scuba Diving Chem Matters Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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