

1. How many grams of NaCl are contained in 700. mL of a 0.250 M solution of sodium chloride?

$$0.250 \frac{\text{mol NaCl}}{\text{L}} \times 0.700 \text{ L} \times \frac{58.5 \text{ g NaCl}}{\text{mol}} = 10.2 \text{ g NaCl}$$

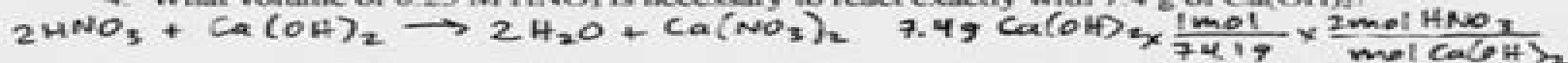
2. The oxidation state of chromium in K_2CrO_4 is:

+6

3. The oxidation state of sulfur in the sulfate ion is

+6

4. What volume of 0.25 M HNO_3 is necessary to react exactly with 7.4 g of Ca(OH)_2 ?



5. Describe how you would prepare 250. mL of a 0.354 M sodium sulfate solution using solid sodium sulfate crystals.

$$0.354 \frac{\text{mol Na}_2\text{SO}_4}{\text{L}} \times 0.250 \text{ L} \times \frac{142 \text{ g Na}_2\text{SO}_4}{\text{mol}} = 12.6 \text{ g}$$

Transfer 12.6 g Na_2SO_4 to a 250 mL vol. flask.

Add distilled water, about 200 mL. Swirl until all solute has dissolved, make up to the mark with distilled water.

$$\times \frac{1 \text{ Liter}}{0.25 \text{ mol HNO}_3} = \boxed{0.80 \text{ L}}$$

- a. What volume, in mL, of 12.1 M hydrochloric acid would be required to make 5.00 L of 2.00 M hydrochloric acid solution?

$$\text{Use } M_1V_1 = M_2V_2$$

$$V_1 = \frac{M_2V_2}{M_1} = \frac{(2.00 \text{ M})(5.00 \text{ L})}{12.1 \text{ M}} = 0.826 \text{ L or } \boxed{826 \text{ mL}}$$

6. Predict the products of the following reactions. Where a reaction occurs, write the net ionic equation. If no reaction occurs, write "N.R."

- a. iron (III) sulfate solution and strontium chloride solution



- b. lead (II) acetate solution and lithium iodide solution



- c. aluminum nitrate solution and calcium bromide solution

N.R.

- d. ammonium hydroxide solution and cobalt (II) nitrate powder



Gas Variables Worksheet Answers

Karen M. Perrin, Perrin



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