



A titrimetric analysis requires the careful addition of titrant.

Experiment 9

A Volumetric Analysis

- To prepare and standardize a sodium hydroxide solution
- To determine the molar concentration of a strong acid

OBJECTIVES

The following techniques are used in the Experimental Procedure:

TECHNIQUES



A chemical analysis that is performed primarily with the aid of volumetric glassware (e.g., pipets, burets, volumetric flasks) is called a **volumetric analysis**. For a volumetric analysis procedure, a known quantity or a carefully measured amount of one substance reacts with a to-be-determined amount of another substance with the reaction occurring in aqueous solution. The volumes of all solutions are carefully measured with volumetric glassware.

The known amount of the substance for an analysis is generally measured and available in two ways:

1. As a **primary standard**—An accurate mass (and thus, moles) of a solid substance is measured on a balance, dissolved in water, and then reacted with the substance being analyzed.
2. As a **standard solution**—A measured number of moles of substance is present in a measured volume of solution, generally expressed as the molar concentration (or molarity) of the substance. A measured volume of the standard solution then reacts with the substance being analyzed.

The reaction of the known substance with the substance to be analyzed, occurring in aqueous solution, is generally conducted by a titration procedure.

The titration procedure requires a buret to dispense a liquid, called the **titrant**, into a flask containing the **analyte** (Figure 9.1a, page 134). For the acid–base titration studied in Part B of this experiment, the titrant is a standard solution of sodium hydroxide and the analyte is an acid.

INTRODUCTION

Primary standard: a substance that has a known high degree of purity, a relatively large molar mass, is nonhygroscopic, and reacts in a predictable way

Standard solution: a solution having a very well known concentration of a solute

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Table of Contents Lab Manual Beran Hein

1. Understanding the eBook Lab Manual Beran Hein
 - The Rise of Digital Reading Lab Manual Beran Hein
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Manual Beran Hein
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Manual Beran Hein
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Manual Beran Hein
 - Personalized Recommendations
 - Lab Manual Beran Hein User Reviews and Ratings

- Lab Manual Beran Hein and Bestseller Lists
- 5. Accessing Lab Manual Beran Hein Free and Paid eBooks
 - Lab Manual Beran Hein Public Domain eBooks
 - Lab Manual Beran Hein eBook Subscription Services
 - Lab Manual Beran Hein Budget-Friendly Options
- 6. Navigating Lab Manual Beran Hein eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Manual Beran Hein Compatibility with Devices
 - Lab Manual Beran Hein Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Manual Beran Hein
 - Highlighting and Note-Taking Lab Manual Beran Hein
 - Interactive Elements Lab Manual Beran Hein
- 8. Staying Engaged with Lab Manual Beran Hein
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Manual Beran Hein
- 9. Balancing eBooks and Physical Books Lab Manual Beran Hein
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Manual Beran Hein
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab Manual Beran Hein
 - Setting Reading Goals Lab Manual Beran Hein
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lab Manual Beran Hein
 - Fact-Checking eBook Content of Lab Manual Beran Hein
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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