

LAB #5 - LIMITING REAGENT**Objective**

- Use stoichiometry to determine the limiting reactant.
- Calculate the theoretical yield.
- Calculate the percent yield of a reaction.

Introduction

In lecture you have learned to read chemical equations and evaluate the mol to mol ratios of reactants and products involved in a chemical reaction. In laboratory experiments it is difficult to measure out chemicals in the exact ratio necessary for the chemical reaction. For time and speed reasons, the reaction mixtures in lab will usually have a limiting and an excess reactant.

Limiting reagent (also called limiting reactant) problems use stoichiometry to determine the theoretical yield for a chemical reaction. The *limiting reactant* will be completely consumed in the reaction and limits the amount of product you can make. The limiting reactant also determines the amount of product you can make (the *theoretical yield*). The reactant that is left over after the reaction is complete is called the *excess reactant*.

Example 1:

Consider the process it takes to make a ham sandwich. You need two slices of bread and one piece of ham to make each sandwich. How many complete sandwiches could you make if you had eighteen slices of bread and six slices of ham? Let's set this up like a chemical equation where the ham and bread are our reactants and the sandwich is our product:



Now we can use stoichiometry to determine the amount of sandwiches we could make if we used all of our reactants.

$$18 \text{ slices of bread} \times \frac{1 \text{ sandwich}}{2 \text{ pieces of bread}} = 9 \text{ sandwiches}$$

$$6 \text{ slices of ham} \times \frac{1 \text{ sandwich}}{1 \text{ slice of ham}} = 6 \text{ sandwiches}$$

Laboratory Manual For Introductory Chemistry Limiting Reagent

Thomas Griffiths



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Table of Contents Laboratory Manual For Introductory Chemistry Limiting Reagent

1. Understanding the eBook Laboratory Manual For Introductory Chemistry Limiting Reagent
 - The Rise of Digital Reading Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Advantages of eBooks Over Traditional Books
2. Identifying Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Laboratory Manual For Introductory Chemistry Limiting Reagent
 - User-Friendly Interface
4. Exploring eBook Recommendations from Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Personalized Recommendations
 - Laboratory Manual For Introductory Chemistry Limiting Reagent User Reviews and Ratings
 - Laboratory Manual For Introductory Chemistry Limiting Reagent and Bestseller Lists
5. Accessing Laboratory Manual For Introductory Chemistry Limiting Reagent Free and Paid eBooks
 - Laboratory Manual For Introductory Chemistry Limiting Reagent Public Domain eBooks
 - Laboratory Manual For Introductory Chemistry Limiting Reagent eBook Subscription Services
 - Laboratory Manual For Introductory Chemistry Limiting Reagent Budget-Friendly Options

6. Navigating Laboratory Manual For Introductory Chemistry Limiting Reagent eBook Formats
 - ePub, PDF, MOBI, and More
 - Laboratory Manual For Introductory Chemistry Limiting Reagent Compatibility with Devices
 - Laboratory Manual For Introductory Chemistry Limiting Reagent Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Highlighting and Note-Taking Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Interactive Elements Laboratory Manual For Introductory Chemistry Limiting Reagent
8. Staying Engaged with Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Laboratory Manual For Introductory Chemistry Limiting Reagent
9. Balancing eBooks and Physical Books Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Laboratory Manual For Introductory Chemistry Limiting Reagent
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Setting Reading Goals Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Fact-Checking eBook Content of Laboratory Manual For Introductory Chemistry Limiting Reagent
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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