

LAB #5 - LIMITING REAGENT**Objective**

- Use stoichiometry to determine the limiting reactant.
- Calculate the theoretical yield.
- Calculate the percent yield of a reaction.

Introduction

In lecture you have learned to read chemical equations and evaluate the mol to mol ratios of reactants and products involved in a chemical reaction. In laboratory experiments it is difficult to measure out chemicals in the exact ratio necessary for the chemical reaction. For time and speed reasons, the reaction mixtures in lab will usually have a limiting and an excess reactant.

Limiting reagent (also called limiting reactant) problems use stoichiometry to determine the theoretical yield for a chemical reaction. The *limiting reactant* will be completely consumed in the reaction and limits the amount of product you can make. The limiting reactant also determines the amount of product you can make (the *theoretical yield*). The reactant that is left over after the reaction is complete is called the *excess reactant*.

Example 1:

Consider the process it takes to make a ham sandwich. You need two slices of bread and one piece of ham to make each sandwich. How many complete sandwiches could you make if you had eighteen slices of bread and six slices of ham? Let's set this up like a chemical equation where the ham and bread are our reactants and the sandwich is our product:



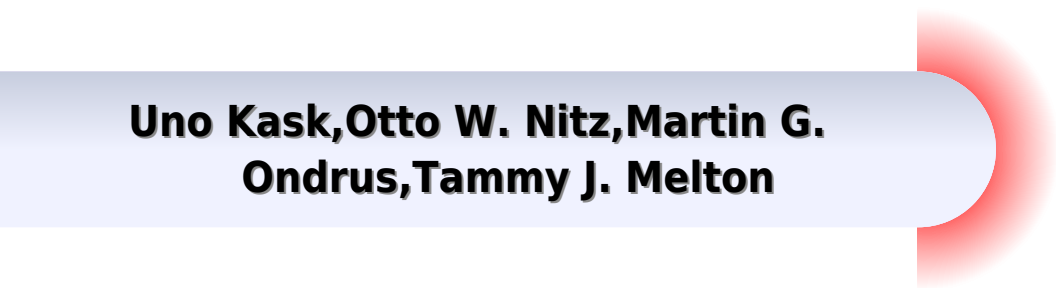
Now we can use stoichiometry to determine the amount of sandwiches we could make if we used all of our reactants.

$$18 \text{ slices of bread} \times \frac{1 \text{ sandwich}}{2 \text{ pieces of bread}} = 9 \text{ sandwiches}$$

$$6 \text{ slices of ham} \times \frac{1 \text{ sandwich}}{1 \text{ slice of ham}} = 6 \text{ sandwiches}$$

Laboratory Manual For Introductory Chemistry Limiting Reagent

**Uno Kask, Otto W. Nitz, Martin G.
Ondrus, Tammy J. Melton**



Laboratory Manual For Introductory Chemistry Limiting Reagent:

Laboratory Manual For Introductory Chemistry Limiting Reagent Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Laboratory Manual For Introductory Chemistry Limiting Reagent**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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