#### AP\* CHEMISTRY 2008 SCORING GUIDELINES

#### Charaction 2

Answer the following appetions relating to praying the profession

In the first of two experiments, a student is assigned the task of determining the number of motes of water in one mote of MgCL, in H<sub>o</sub>O. The student collects the data shows in the following table.

Mass of empty container	22.347 g
Initial mass of sample and container-	25.825 g
Mass of sample and sometime after first heating	23.982 g
Mass of sample and container after second heating	23,978 g
Mass of sample and container after third heating	23.977.8

(x) Explain why the student can correctly conclude that the hydrone was heated a sufficient number of times in the experiment.

No additional mass was lost during the third bearing, indicating that all the water of hydration had been driven off.

One point is named for the cornect explanation.

- (b). Use the data above to
  - 63 cultivities the total number of motor of winer live when the sample was beased, and

mass of 
$$H_2O$$
 lost =  $25.825 - 23.977 = 1.848 g$   
 $OR$   
 $25.825 - 23.906 = 1.849 g$   
 $1.848 g H_2O \times \frac{1 \mod H_2O}{18.02 g H_2O} = 0.1028 \text{ and } H_2O$ 

One point is assed for calculating the correct number of rucles of water.

(iii) desermine the formula of the hydrated compound.

mass of unhydrous 
$$MgCl_2 = 23.977 - 22.347 = 1.630 g$$

$$0.830 \pm MgCl_2 \times \frac{1 \bmod MgCl_2}{95.30 \pm MgCl_2} = 0.01712 \bmod MgCl_3$$

$$\frac{0.1026 \text{ mod H}_2\text{O}}{0.00712 \text{ mod MgCI}_2} = 5.993 = 6 \text{ mod H}_2\text{O per mod MgCI}_2$$

Formula is MgCl<sub>3</sub> -6H<sub>2</sub>O

One point is earned for calculating the correct number of incise of anhydrous MgCl<sub>2</sub>.

One point is carned for writing the correct formula (with supporting calculations).

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**Yi-Tong Ma** 

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