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EDUCATION

VHEMBE EAST DISTRICT

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NATIONAL SENIOR SCHOOL CERTIFICATE

TERM 1 COMBINED TEST

2022

MARKS: 30

TIME: 45 MINUTES

DIRECTIONS AND INFORMATION

- The paper consists of papers including cover
- Read all questions and/or questions paper carefully
- Answer all questions
- Write neatly and clearly

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
















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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution $4\text{H}_2\text{O} + 5\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_3^- \rightarrow 2\text{MnO}_2 + \text{Cr}_3\text{O}_3 + 2\text{NO}_2 + 2\text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}_3\text{Te} + 10\text{ReO}_4 + 11\text{PbO}_2 + \text{IO}_2 + \text{Pb}_2 + 12\text{H}^+$ as $\text{NO}_3^- + \text{IO}_3^- + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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