

P1.1 A gas at 20°C may be *rarefied* if it contains less than 10^{12} molecules per mm^3 . If Avogadro's number is $6.023\text{E}23$ molecules per mole, what air pressure does this represent?

Solution: The mass of one molecule of air may be computed as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.023\text{E}23 \text{ molecules/g}\cdot\text{mol}} = 4.81\text{E}-23 \text{ g}$$

Then the density of air containing 10^{12} molecules per mm^3 is, in SI units,

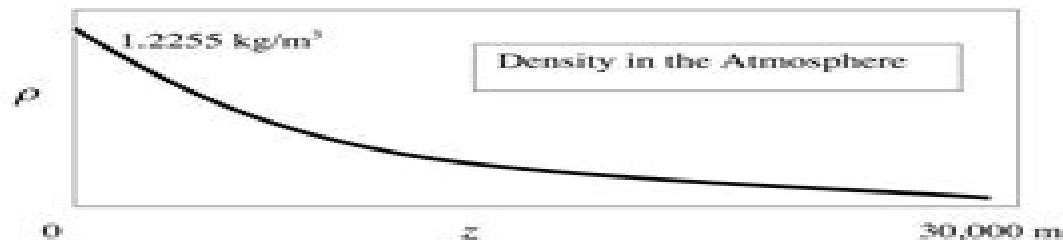
$$\begin{aligned} \rho &= \left(10^{12} \frac{\text{molecules}}{\text{mm}^3} \right) \left(4.81\text{E}-23 \frac{\text{g}}{\text{molecule}} \right) \\ &= 4.81\text{E}-11 \frac{\text{g}}{\text{mm}^3} = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \end{aligned}$$

Finally, from the perfect gas law, Eq. (1.13), at 20°C = 293 K, we obtain the pressure:

$$p = \rho RT = \left(4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \right) \left(287 \frac{\text{m}^2}{\text{s}^2\cdot\text{K}} \right) (293 \text{ K}) = 4.0 \text{ Pa} \quad \text{Ans.}$$

P1.2 Table A.6 lists the density of the standard atmosphere as a function of altitude. Use these values to estimate, crudely, say, within a factor of 2, the number of molecules of air in the entire atmosphere of the earth.

Solution: Make a plot of density ρ versus altitude z in the atmosphere, from Table A.6:



This writer's approximation: The curve is approximately an exponential, $\rho = \rho_0 \exp(-b z)$, with b approximately equal to 0.00011 per meter. Integrate this over the entire atmosphere, with the radius of the earth equal to 6377 km:

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