

Electron Configuration Lab

Name _____ KEY _____ Period _____ Date _____

Directions: To understand that electrons fill up the LOWEST energy orbitals first, but that they spread out as much as possible in orbitals.

Part 1: Questions

1. Write the Aufbau Principle.

Electrons will fill the lowest energy orbitals first.

2. Which "rule" or "principle" states that electrons spread out as much as possible within a sublevel before they pair up in the same orbital? **Hund's Rule**

3. In this lab, one cup represents an orbital in the atom. How many cups would be needed to represent the "2p" sublevel?

three (3)

4. How many cups would be needed to represent an atom that has the 1s, 2s, 2p, 3s, 3p sublevels _____?

ten (10)

5. In this lab, beans represent the electrons. How many beans can go into a single cup?

two (2)

Part 2:

Part 2:

Electron Configuration	Total # of Electrons	Element Name	Teacher's Initials
$1s^2 2s^2 2p^6 3s^2 3p^3$	15	Phosphorus	

Part 3:

Station	Electron Configuration	Total # of Electrons
Station 1	$1s^2 2s^2 2p^6$	10
Station 2	$1s^2 2s^2 2p^6$	8
Station 3	$1s^2 2s^2 2p^6 3s^2 3p^2 4s^2$	20
Station 4	$1s^2 2s^2 2p^6 3s^2 3p^6$	18
Station 5	$1s^2 2s^2 2p^3$	7
Station 6	$1s^2 2s^2$	4
Station 7	$1s^2 2s^2 2p^3$	6
Station 8	$1s^2 2s^2$	4
Station 9	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	18
Station 10	$1s^2 2s^2 2p^3$	7

Part 4: Questions

1. Using a Periodic Table, identify the elements at each station and complete the table to the right.

2. Which orbital is lower in energy -- 4s or 3d?

4s

3. Write the electron configuration for Zinc.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$

4. Write the electron configuration for Xenon.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6$

5. Write the electron configuration for Bismuth.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^3$

Station	Element Name
Station 1	Neon (Ne)
Station 2	Oxygen (O)
Station 3	Calcium (Ca)
Station 4	Argon (Ar)
Station 5	Boron (B)
Station 6	Beryllium (Be)
Station 7	Carbon (C)
Station 8	Lithium (Li)
Station 9	Potassium (K)
Station 10	Nitrogen (N)

Electron Configuration Lab

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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations Exploring General Chemistry

in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **2024-25 NVS Lab Attendant/Assistant Solved Papers** YCT Expert Team , 2024 25 NVS Lab

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