Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- nitrogen 1s²2s²2p³
- copper 1s² 2s² 2p⁶ 3s² 3p⁶ 4s¹ 3d¹⁰
- chlorine 1s² 2s² 2p⁶ 3s² 3p⁵
- calcium 1s² 2s² 2p⁶ 3s² 3p⁶ 4s²
- 5) lead 1s² 2s² 2p⁶ 3s²3p⁶4s² 3d¹⁶ 4p⁶ 5s² 4d¹⁸ 5p⁵ 6s² 4f¹⁴ 5d¹⁶ 6p²

In the space below, write the <u>noble gas notation</u> electron configurations of the following elements:

- 6) barium [Xe] 6s2
- 7) gold [Xe] 4f¹⁴5d¹⁰6s¹
- B) platinum [Xe] 4f¹⁴ 5d⁹ 6s¹
- yenon [Kr] 4d¹⁰ 5s² p⁶
- 10) calcium [Ar] 4s2

Determine what elements are denoted by the following electron configurations:

- 11) 1s²2s²2p⁶3s²3p³ phosphorous
- 12) 1s²2s²2p⁶3s²3p⁶4s²3d¹⁰4p⁶5s¹4d⁷ ruthenium
- 13) [Kr] 4d¹⁰5s²p⁵ lodine
- 14)[Xe] 4f145d106s2 mercury
- 15)[Rn] 7s² radium

Determine which of the following electron configurations are not correct:

- 16)1s²2s²2p⁶3s²3p⁵4s² 3p⁵ should be a 6
- 17)1s³2s²2p⁶3s²2p⁵ 1s only holds 2 and there aren't two "2p" orbitals
- 18)[Cs] 5f¹⁴6d¹7s² Cs is not a noble gas
- 19)[Kr] 4d105s2p4 Correct
- 20)[Ne] for Neon Can't use a noble gas to shortcut itself

Electron Configurations Packet Answers

Ying Liu



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