

Assignment #3 – Compound Names and Formulas
Single and Multi-valent ions

A. Write the correct formula for the following compounds.

1. gallium oxide Ga_2O_3
2. mercury (I) chloride $HgCl$
3. tin (IV) bromide $SnBr_4$
4. indium oxide In_2O_3
5. cadmium sulphide CdS
6. magnesium oxide MgO
7. potassium nitride K_3N
8. iron (III) sulphide Fe_2S_3
9. gold (III) bromide $AuBr_3$
10. zinc oxide ZnO

11. silicon oxide SiO_2
12. aluminum bromide $AlBr_3$
13. tin (II) sulphide SnS
14. germanium bromide $GeBr_4$
15. nickel (II) oxide NiO
16. strontium astatide $SrAt_2$
17. aluminum oxide Al_2O_3
18. copper (II) oxide CuO
19. lawrencium fluoride LrF_3
20. actinium oxide Ac_2O_3

B. Name the following compounds correctly. Only include the Roman Numeral in the name if the first element has 2 or more combining capacities.

1. $CaBr_2$ Calcium Bromide
2. K_2O Potassium Oxide
3. Fe_2O_3 Iron (III) Oxide
4. ZnF_2 Zinc Fluoride
5. $CuCl$ Copper (I) Chloride
6. Ni_2O_3 Nickel (III) Oxide
7. $AuCl$ Gold (I) Chloride
8. $HgBr$ Mercury (I) Bromide

9. $AsCl_5$ Arsenic (V) Chloride
10. Eu_2O_3 Europium (III) Oxide
11. $NdBr_3$ Neodymium Bromide
12. Cu_3N_2 Copper (II) Nitride
13. CoS Cobalt (II) Sulphide
14. ZrO_2 Zirconium Oxide
15. Li_2O Lithium Oxide
16. NaH Sodium Hydride

Naming Ionic Compounds If8766 Answer

Thomas Griffiths



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