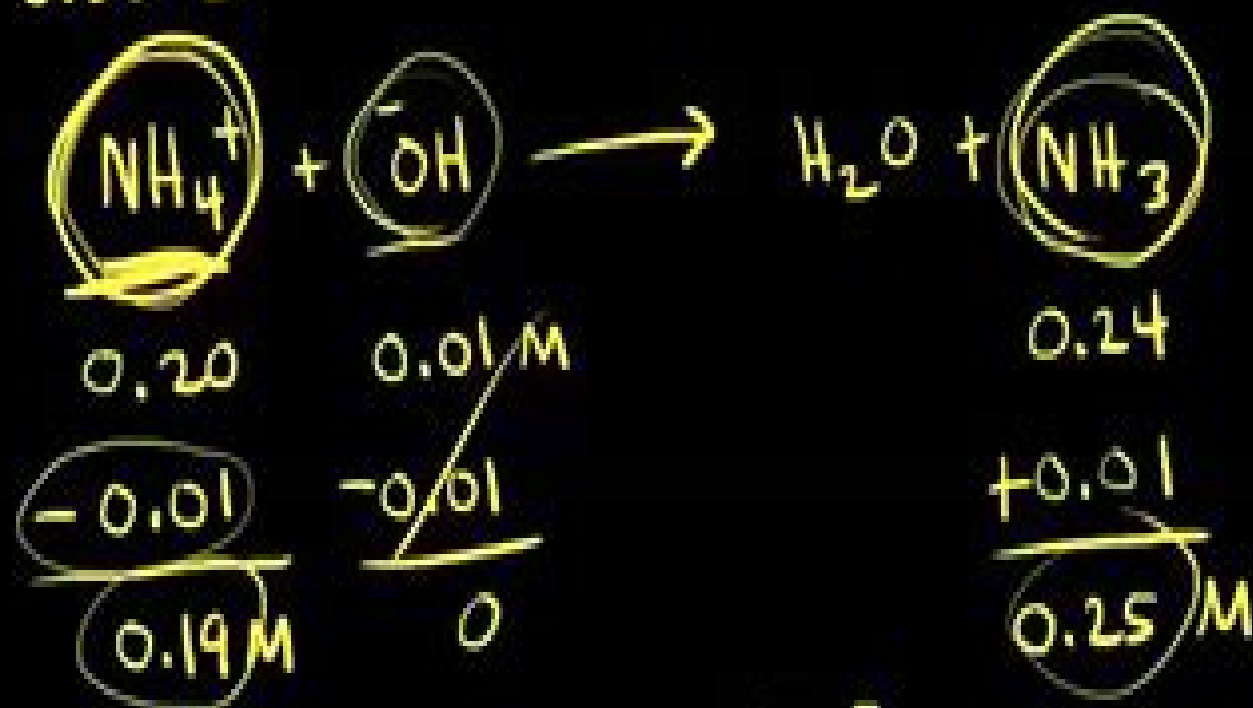


If 0.005 mol of NaOH is added to 0.50 L of the buffer solution (0.24 M NH_3 and 0.20 M NH_4Cl), what is the resulting pH?

$$\frac{0.005 \text{ mol}}{0.50 \text{ L}} = 0.01 \text{ M } [\text{OH}^-]$$



$$\text{pH} = 9.25 + \log \frac{0.25}{0.19}$$

$$\text{pH} = \boxed{9.37}$$

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AW Chickering



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