

## Electron Energy and Light

How does light reveal the behavior of electrons in an atom?

### Why?

From fireworks to stars, the color of light is useful in finding out what's in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be identified by heating or burning the materials and examining the color(s) of light given off in the form of bright-line spectra.

### Model 1 – White Light



- Trace the arrows in Model 1 and shade in the table with colored pencils where appropriate.
- What happens to white light when it passes through a prism?  
The light is split into colors of continuous visible spectrum.
- Why are the color labels in the table in Model 1 plural (i.e., "Reds" rather than "Red")?  
There are several different shades of all the colors that are in the same wavelength range.
- Do all colors of light travel at the same speed?  
Yes.
- Do all colors of light have the same energy? If no, which colors have the highest energy and the least energy, respectively?  
No, all colors do not have same energy. The violet and blues have most energy and red and oranges have the least energy.
- Consider the light illustrated in Model 1.
  - Which color corresponds to the longest wavelength?  
The reds.
  - Which color corresponds to the shortest wavelength?  
The violets.
  - Write a sentence that describes the relationship between wavelength and energy of light.  
The energy of light increases as the wavelength gets shorter.



# Electron Energy And Light Pogil Worksheet Answers

**EW Minium**



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