Name:	Date:	Period:
Electron configuration		
The area in which there is 90% probability of fine orbital can holdtwo electrons.	ding an electron is called an	orbital An
There orbitals can be represented in the following	g way:	
Aufba	au Diagram	
Increasing 3s 3px 3py 3pz  Energy 2px 2py 2pz  2s  1s		
Each box represents an orbital. Electrons are desclockwise or counterclockwise spins.	ignated by arrows pointing up	or down which represent
To determine where the electrons of a particular e	element are located you use the	following rules:
RULE 1: An electron will always be found in t	the lowest energy orbital ava-	ilable that has room.
Therefore the electron configuration of Hydrogen	n is: This is designated a	s: 1s <sup>1</sup>
RULE 2: An orbital may hold a maximum of 2 RULE 3: The Pauli Exclusion Principle – Two		will have opposite spins.
Therefore the electron configuration of Helium is		s: 1s <sup>2</sup>
Draw an orbital diagram and write the electron     Li	Be 1s <sup>2</sup> 2s	ryllium, and Boron.
-		

# **Electron Configuration Packet Key**

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