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$$\begin{array}{l} 1 \text{ N}_2 + 2 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \\ \text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{ SO}_2 + 5 \text{ NF}_3 + 2 \text{ SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + \text{CO}_2 \\ 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ SO}_2 + 6 \text{ NaBr} + 3 \text{ SNO}_2 + 3 \text{ NF}_3 + 2 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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assuming $\text{CeCl}_2 \cdot 2\text{H}_2\text{O}$ is available in excess

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 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_4$
 br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_3$
 $\text{NaBr} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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