



# Julius Caesar Study Guide Questions Answers Act 1 Scene

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web using algebra the moles was founded to be 0.00352 moles of butane in order to find the molarity the mass of the butane should be divided by the moles after further calculations the molarity of the butane gas was calculated to be 56.81 grams moles this lab is done with tons of sources of error

web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp v 1 n 1 v 2 n 2 0.0896 l n 22.4 l 1 mole n 0.0040 moles 12 experimental molar mass g mole

of butane 0.23 g 0.0040 moles 58 g mole 13 the molecular formula of butane is  $C_4H_{10}$

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