

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 4****SHORT ANSWER** Answer the following questions in the space provided.

1. b In metals, the valence electrons are considered to be
(a) attached to particular positive ions. (c) immobile.
(b) shared by all surrounding atoms. (d) involved in covalent bonds.
2. a The fact that metals are malleable and ionic crystals are brittle is best explained in terms of their
(a) chemical bonds. (c) enthalpies of vaporization.
(b) London forces. (d) polarity.
3. d As light strikes the surface of a metal, the electrons in the electron sea
(a) allow the light to pass through.
(b) become attached to particular positive ions.
(c) fall to lower energy levels.
(d) absorb and re-emit the light.
4. d Mobile electrons in the metallic bond are responsible for
(a) luster. (c) electrical conductivity.
(b) thermal conductivity. (d) All of the above.
5. a In general, the strength of the metallic bond _____ moving from left to right on any row of the periodic table.
(a) increases (c) remains the same
(b) decreases (d) varies
6. c When a metal is drawn into a wire, the metallic bonds
(a) break easily. (c) do not break.
(b) break with difficulty. (d) become ionic bonds.
7. Use the concept of electron configurations to explain why the number of valence electrons in metals tends to be less than the number in most nonmetals.

Most metals have their outer electrons in s orbitals, while nonmetals have their
outer electrons in p orbitals.

Introduction To Chemical Bonding Review Answers

S Marginson



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