

The Identity of an Insoluble Precipitate

Date: _____
Course: _____
Section: _____
Instructor: _____

Student name: _____
Team members: _____

Prelaboratory assignment

- The following reactions, shown in unbalanced equations, are pertinent to this experiment. Balance the equations.
 - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{Ba}(\text{NH}_2\text{SO}_3)_2 + \text{HNO}_3$
 - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{BaSO}_4 + \text{NH}_4\text{NO}_3 + \text{HNO}_3$
 - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{Ba}(\text{NH}_2)_2 + \text{H}_2\text{SO}_4 + \text{HNO}_3$
- When a solution containing 1.4 g of $\text{Ba}(\text{NO}_3)_2$ and 2.4 g of $\text{NH}_2\text{SO}_3\text{H}$ is boiled, a precipitate forms. This precipitate may be $\text{Ba}(\text{NH}_2\text{SO}_3)_2$, BaSO_4 , or $\text{Ba}(\text{NH}_2)_2$. You can find these products in the reactions whose equations you balanced in the first question.
 - Calculate the number of moles of each reactant.

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