

 **K_{sp} and ICE Problems Explained**

Key: K_{sp} = solubility product
 subscripted values

- a. Write out the solubility product value after writing the dissolving solid's equation in the equation:



Initial	K_{sp}	Change	Equilibrium
PbCl_2	1.7×10^{-2}	$-x$	x
Pb^{2+}		$+x$	x
Cl^-		$+2x$	$2x$

- b. Write out the solubility product value if you have a solid for something with a coefficient in the equation:



Initial	K_{sp}	Change	Equilibrium
Al(OH)_3	3×10^{-34}	$-x$	x
Al^{3+}		$+x$	x
OH^-		$+3x$	$3x$

- c. Use what you learned in question b when you're asked to calculate solubility and K_{sp} . Write out the value that you're looking for after equilibrium:



If you're given solubility, you can calculate K_{sp} . If you're given K_{sp} , you can calculate solubility. Write out the value that you're looking for after equilibrium:

Initial	K_{sp}	Change	Equilibrium	K_{sp}	Equilibrium
PbCl_2	1.7×10^{-2}	$-x$	x	1.7×10^{-2}	x
Pb^{2+}		$+x$	x		x
Cl^-		$+2x$	$2x$		$2x$

Ice Practice Problems Answer Key

Lyn English, Tamara Moore



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