

NOTES - Unit 4: Bonding

PART 1: Ionic and Covalent Bonding

CHEMICAL BONDS

- A chemical bond is an attraction between 2 atoms or ions.
- Bonding occurs because it _____ of the system.
- Three broad classifications (general rule of thumb)
 - Ionic (_____ + _____)
 - Covalent (_____ + _____)
 - Metallic (_____ + _____)
- _____: the power of an atom in a molecule to attract electrons to itself.
 - The greater an atom's electronegativity, the greater its ability to attract _____.
 - In a compound, the element with the greater electronegativity will be the more _____ species.
 - The element with _____ electronegativity is sometimes referred to as the more " _____ " element. This will become the more positive species in the bond.
- _____ bonds exist between atoms with low electronegativities & those with high electronegativities. In general, between metals and nonmetals.
- _____ bonds exist between dissimilar atoms with different electronegativities.
- _____ bonds exist between identical nonmetallic atoms or nonmetallic atoms with similar electronegativities.
- Atoms react with each other in chemical reactions in a quest to have complete outer electron energy levels like the _____.
 - Octet Rule: Atoms tend to gain, lose, or share electrons until they are surrounded by _____ valence electrons.
 - Duet Rule: _____ and _____ want to have 2 electrons. Recall that the first energy level only holds two electrons.

IONIC BONDING

- Different texts report different "cutoff" values, but according to IB, _____ for ionic bonds.
- The greater the difference in electronegativity between the atoms, the more _____ the bond.
- Ionic bonds form between metals on the _____ of the periodic table and elements in groups _____ (or polyatomic ions attracted to other ions)
 - Note: group 4 elements (C and Si) do not form ionic bonds. They tend to form giant molecular structures (network solids) or simple molecules.
- Example: NaCl
 - Sodium (Na) has 11 electrons (but only one valence electron): _____
 - By losing this electron Na becomes Na^+ : _____ or _____ which has a full outer shell.
 - Chlorine (Cl) has 17 electrons (seven valence electrons): _____

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DP Hallahan



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Table of Contents Ib Topics 4 And 14 Ap Chapters Zumdahl

1. Understanding the eBook Ib Topics 4 And 14 Ap Chapters Zumdahl
 - The Rise of Digital Reading Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Advantages of eBooks Over Traditional Books
2. Identifying Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ib Topics 4 And 14 Ap Chapters Zumdahl
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Personalized Recommendations
 - Ib Topics 4 And 14 Ap Chapters Zumdahl User Reviews and Ratings

- Ib Topics 4 And 14 Ap Chapters Zumdahl and Bestseller Lists
- 5. Accessing Ib Topics 4 And 14 Ap Chapters Zumdahl Free and Paid eBooks
 - Ib Topics 4 And 14 Ap Chapters Zumdahl Public Domain eBooks
 - Ib Topics 4 And 14 Ap Chapters Zumdahl eBook Subscription Services
 - Ib Topics 4 And 14 Ap Chapters Zumdahl Budget-Friendly Options
- 6. Navigating Ib Topics 4 And 14 Ap Chapters Zumdahl eBook Formats
 - ePub, PDF, MOBI, and More
 - Ib Topics 4 And 14 Ap Chapters Zumdahl Compatibility with Devices
 - Ib Topics 4 And 14 Ap Chapters Zumdahl Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Highlighting and Note-Taking Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Interactive Elements Ib Topics 4 And 14 Ap Chapters Zumdahl
- 8. Staying Engaged with Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ib Topics 4 And 14 Ap Chapters Zumdahl
- 9. Balancing eBooks and Physical Books Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ib Topics 4 And 14 Ap Chapters Zumdahl
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Setting Reading Goals Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Fact-Checking eBook Content of Ib Topics 4 And 14 Ap Chapters Zumdahl
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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