

13. Which type of bond is formed when a Lewis acid reacts with a Lewis base?

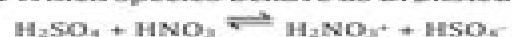
- A. Covalent
- B. Dipole-dipole
- C. Double
- D. Hydrogen

14. What is the order of increasing acidity of the following acids?

Acid	K_a	Acid	pK_a
chloroethanoic	1.3×10^{-3}	hydrogen fluoride	3.3
ethanoic	1.7×10^{-5}	hydrogen cyanide	9.3

- A. chloroethanoic < ethanoic < hydrogen fluoride < hydrogen cyanide
- B. ethanoic < chloroethanoic < hydrogen fluoride < hydrogen cyanide
- C. chloroethanoic < ethanoic < hydrogen cyanide < hydrogen fluoride
- D. hydrogen cyanide < ethanoic < hydrogen fluoride < chloroethanoic

15. Which species behave as Brønsted-Lowry bases in the following reaction?



- A. HNO_3 and HSO_4^-
- B. HNO_3 and H_2NO_3^+
- C. H_2SO_4 and HSO_4^-
- D. H_2NO_3^+ and HSO_4^-

16. Which is a conjugate Brønsted-Lowry acid-base pair?



- A. $\text{CH}_3\text{COO}^- / \text{H}_3\text{O}^+$
- B. $\text{H}_2\text{O} / \text{CH}_3\text{COO}^-$
- C. $\text{H}_2\text{O} / \text{H}_3\text{O}^+$
- D. $\text{CH}_3\text{COOH} / \text{H}_2\text{O}$

17. Which of the following gases does **not** result in acid deposition?

- A. CO_2
- B. NO_2
- C. NO
- D. SO_2

18. What occurs when solid sodium hydrogen carbonate reacts with aqueous sulfuric acid?

- A. Bubbles of sulfur dioxide form.
- B. Bubbles of both hydrogen and carbon dioxide form.
- C. Bubbles of hydrogen form.
- D. Bubbles of carbon dioxide form.

19. A student carried out a titration to determine the concentration of an acid and found that his value had good precision but poor accuracy. Which process explains this outcome?

- A. Consistently overshooting the volume of solution from the burette into the flask.
- B. Collection of insufficient titration data.
- C. Reading the meniscus in the burette at a different angle each time.
- D. Forgetting to rinse the flask after one of the titrations.

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