

Name: _____ Date: _____ Period: _____

Midterm Exam Study Guide

d. Isotopes

- i. Isotopes differ in their _____ and number of _____

e. Average atomic mass Formula:

f. Ions

- i. _____ have lost electrons and have a _____ charge
ii. _____ have gained electrons and have a _____ charge

g. Group Numbers be able to locate them

- i. Alkali Metals
ii. Alkaline earth Metals
iii. Transition Metals
iv. Inner Transition Metals
v. Halogens
vi. Noble Gases

h. Periodic Trends Give your notes for reference (be able to for each of these)

- i. Atomic Radius/Size
ii. Ionic Radius/Size
iii. Ionization Energy
iv. Electronegativity

i. Valence Electrons be able to identify # of valence electrons and charge for each group

- i. Lewis-Dot structures

2. Nuclear Decay and Half Life

- a. ${}^4_2\text{He}$ or ${}^4_2\text{He}$ is called an _____ particle which has the weakest radiation.
b. ${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$ is called an _____ particle which has moderate radiation and is an electron.
c. ${}^1_0\text{n}$ is called an _____ particle which has the strongest radiation and highest energy.
d. Half-life
i. Definition: _____
ii. Know how to read a graph and calculate half-life

3. Electron Configurations

- a. Be able to draw them/identify them from another student
b. Hund's Rule
c. Pauli Exclusion Principle
d. Aufbau Principle

4. Ionic Compounds

- a. Ionic compounds are made of _____ + _____
b. Their overall charge is _____
c. Know how to use Lewis-Dot Structures to draw bonds between cation and anion

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