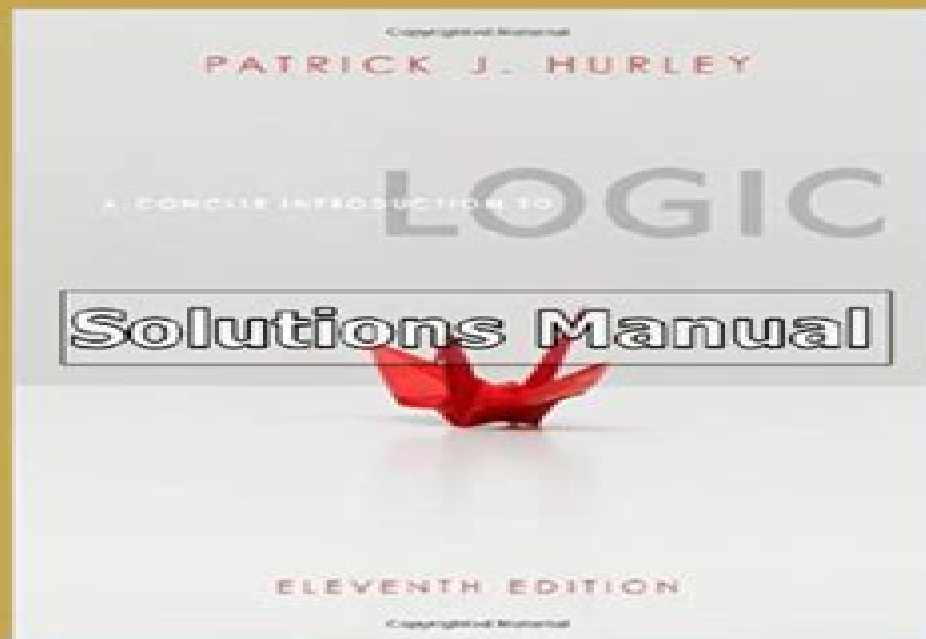


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Introduction To Logic 11th Edition

Richard A. Jones



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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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 $2 \text{NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)} + \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq}$
 $\text{H}_2\text{O (l)} + \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$