

THE MOLE WORKSHEET

1. What is a mole of a substance?

A mole of a substance is 6.02×10^{23} entities (atoms or molecules) of a substance. Therefore, one mole of oxygen has 6.02×10^{23} atoms, which will then have a mass of 16.00 g.

2. What does the term molar mass mean?

A molar mass is the mass of one mole of a substance. Oxygen has a molar mass of 16.00 g/mol (for every mole of oxygen there is, it weighs 16.00 g).

3. What is the mass of the following substances?

(a) 3.7 mol of K

$$\begin{aligned} m &= n \times M \\ &= 3.7 \text{ mol} \times 39.10 \text{ g/mol} \\ &= 144.67 \text{ g} \end{aligned}$$

b) 0.0067 mol of SO_2

$$\begin{aligned} m &= n \times M \\ &= 0.0067 \text{ mol} \times 64.06 \text{ g/mol} \\ &= 0.43 \text{ g} \end{aligned}$$

4. Determine the number of moles of substance represented by each of the following masses.

(a) 5.6 g of He

$$\begin{aligned} n &= m / M \\ &= 5.6 \text{ g} / 4.0 \text{ g/mol} \\ &= 1.4 \text{ mol} \end{aligned}$$

(b) 67 g of MgCl_2

$$\begin{aligned} n &= m / M \\ &= 67 \text{ g} / 95.21 \text{ g/mol} \\ &= 0.703 \text{ mol} \end{aligned}$$

5. How many atoms or molecules are present in each of the following? (Hint: Use Avogadro's constant)

(Note: If you are working with an element, you are calculating for the # of atoms. If you are working with a compound, you are calculating for the # of molecules).

a) 2.0 mol Ne

$$\begin{aligned} \# \text{ atoms} &= n_{\text{Ne}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \\ &= 2.0 \text{ mol} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \\ &= 1.2 \times 10^{24} \text{ atoms} \end{aligned}$$

b) 4.5 g H_2O

$$\begin{aligned} \# \text{ molecules} &= m_{\text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= m_{\text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \frac{m_{\text{H}_2\text{O}}}{M_{\text{H}_2\text{O}}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \frac{4.5 \text{ g}}{18.02 \text{ g/mol}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 1.50 \times 10^{23} \text{ molecules} \end{aligned}$$

c) 0.7 mol CO_2

$$\begin{aligned} \# \text{ molecules} &= n_{\text{CO}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 0.7 \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 4.21 \times 10^{23} \text{ molecules} \end{aligned}$$

Introduction To The Mole Article Answers

DJ Losen



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Table of Contents Introduction To The Mole Article Answers

1. Understanding the eBook Introduction To The Mole Article Answers
 - The Rise of Digital Reading Introduction To The Mole Article Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Introduction To The Mole Article Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Introduction To The Mole Article Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Introduction To The Mole Article Answers
 - Personalized Recommendations
 - Introduction To The Mole Article Answers User Reviews and Ratings
 - Introduction To The Mole Article Answers and Bestseller Lists
5. Accessing Introduction To The Mole Article Answers Free and Paid eBooks
 - Introduction To The Mole Article Answers Public Domain eBooks
 - Introduction To The Mole Article Answers eBook Subscription Services
 - Introduction To The Mole Article Answers Budget-Friendly Options
6. Navigating Introduction To The Mole Article Answers eBook Formats

- ePub, PDF, MOBI, and More
- Introduction To The Mole Article Answers Compatibility with Devices
- Introduction To The Mole Article Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Introduction To The Mole Article Answers
 - Highlighting and Note-Taking Introduction To The Mole Article Answers
 - Interactive Elements Introduction To The Mole Article Answers
- 8. Staying Engaged with Introduction To The Mole Article Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Introduction To The Mole Article Answers
- 9. Balancing eBooks and Physical Books Introduction To The Mole Article Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Introduction To The Mole Article Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Introduction To The Mole Article Answers
 - Setting Reading Goals Introduction To The Mole Article Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Introduction To The Mole Article Answers
 - Fact-Checking eBook Content of Introduction To The Mole Article Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web may 10 2023 l optimisation pour les moteurs de recherche seo est une stratégie de marketing numérique visant à améliorer la visibilité et le classement d un site web sur les pages de résultats des moteurs de recherche serp telles que google bing et yahoo le référencement consiste à optimiser divers éléments du site web tels que le