

(Question 1 continued)

- (d) The cross-sectional area of the tube is $1.3 \times 10^{-3} \text{ m}^2$ and the temperature of air is 300 K. Estimate the number of moles of air in the tube. [2]

$$\begin{aligned}
 PV &= nRT & (b) A &= PY & \text{or } n &= \frac{LA}{Y} \\
 n &= \frac{PV}{RT} & & & & \frac{LA}{Y} \\
 & & & & & n = \frac{(5 \times 10^{-4})(1.3 \times 10^{-3})}{(8.31)(300)} \\
 & & & & & n = 0.026 \text{ mol} \\
 & & & & & \approx 0.03 \text{ mol}
 \end{aligned}$$

but given gives Area = $1.3 \times 10^{-3} \text{ m}^2$

Since $Y = \frac{F}{A} \text{ or } F = YA$

- (e) The equation in (b) may be used to predict the pressure of the air at extremely large values of $\frac{1}{H}$. Suggest why this will be an unreliable estimate of the pressure. [2]

Very large $1/H$ means small volume. Small volume means high pressure.

In ideal gas law, it stated that molecules are not close to each other and hence the forces between them are negligible.

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Table of Contents Ib Physics Sl Paper 3 2013

1. Understanding the eBook Ib Physics Sl Paper 3 2013
 - The Rise of Digital Reading Ib Physics Sl Paper 3 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Ib Physics Sl Paper 3 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ib Physics Sl Paper 3 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Physics Sl Paper 3 2013
 - Personalized Recommendations
 - Ib Physics Sl Paper 3 2013 User Reviews and Ratings
 - Ib Physics Sl Paper 3 2013 and Bestseller Lists
5. Accessing Ib Physics Sl Paper 3 2013 Free and Paid eBooks

- Ib Physics Sl Paper 3 2013 Public Domain eBooks
- Ib Physics Sl Paper 3 2013 eBook Subscription Services
- Ib Physics Sl Paper 3 2013 Budget-Friendly Options
- 6. Navigating Ib Physics Sl Paper 3 2013 eBook Formats
 - ePub, PDF, MOBI, and More
 - Ib Physics Sl Paper 3 2013 Compatibility with Devices
 - Ib Physics Sl Paper 3 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ib Physics Sl Paper 3 2013
 - Highlighting and Note-Taking Ib Physics Sl Paper 3 2013
 - Interactive Elements Ib Physics Sl Paper 3 2013
- 8. Staying Engaged with Ib Physics Sl Paper 3 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ib Physics Sl Paper 3 2013
- 9. Balancing eBooks and Physical Books Ib Physics Sl Paper 3 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ib Physics Sl Paper 3 2013
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ib Physics Sl Paper 3 2013
 - Setting Reading Goals Ib Physics Sl Paper 3 2013
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ib Physics Sl Paper 3 2013
 - Fact-Checking eBook Content of Ib Physics Sl Paper 3 2013
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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