

4. Which of the following substances would have hydrogen bonding as one of its IMF's?



5. Which molecular substance from the following list would have the highest boiling point?

- A. Ne
- B. F_2
- C. H_2
- D. Cl_2

6. When the actual gas pressure is less than the pressure predicted by the ideal gas law, the explanation lies in the fact that the ideal gas law does NOT include a factor for molecular

- A. volume
- B. mass
- C. velocity
- D. attraction
- E. shape

7. 25.0 L of hydrogen gas is collected over water at 40°C . The total pressure of the sample is 750 mmHg. At 40°C , the vapor pressure of water is 55 mmHg. What is the mass of the hydrogen gas?

- A. 1.00 g
- B. 1.50 g
- C. 0.0060 g
- D. 0.0607 g
- E. 1.12 g

8. A 2.00 L sample of nitrogen gas at 27°C and 600 mmHg is heated until it occupies a volume of 3.00 L. If the pressure remains unchanged, the final temperature of the gas is

- A. 40°C
- B. 120°C
- C. 477°C
- D. 677°C
- E. 900°C

9. PI (d) bonding occurs in each of the following EXCEPT

- A. C_2H_2
- B. C_2H_4
- C. C_2H_6
- D. C_2H_2
- E. C_2H_4

10. Determine the shape and polarity of the PF_3 molecule

- A. trigonal pyramidal, polar
- B. trigonal pyramidal, nonpolar
- C. trigonal planar, polar
- D. trigonal planar, nonpolar

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Kathleen Armour



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