

Gases

Section 13.1 The Gas Laws

pages 442–451

Practice Problems

page 443

Assume that the temperature and the amount of gas are constant in the following problems.

1. The volume of a gas at 99.0 kPa is 300.0 mL. If the pressure is increased to 188 kPa, what will be the new volume?

158 mL

$$V_2 = \frac{V_1 P_1}{P_2} = \frac{(300.0 \text{ mL})(99.0 \text{ kPa})}{188 \text{ kPa}} = 158 \text{ mL}$$

2. The pressure of a sample of helium in a 1.00-L container is 0.988 atm. What is the new pressure if the sample is placed in a 2.00-L container?

0.494 atm

$$P_2 = \frac{V_1 P_1}{V_2} = \frac{(1.00 \text{ L})(0.988 \text{ atm})}{2.00 \text{ L}} = 0.494 \text{ atm}$$

3. **Challenge** Air trapped in a cylinder fitted with a piston occupies 145.7 mL at 1.08 atm pressure. What is the new volume when the piston is depressed, increasing the pressure by 25%?

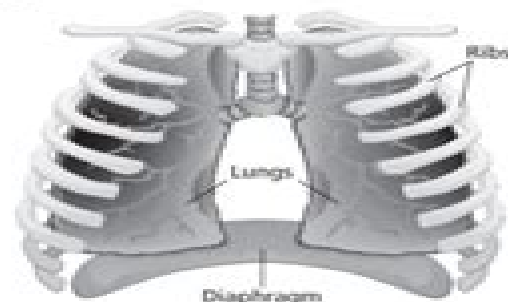
117 mL

$$P_2 = 1.08 \text{ atm} + (1.08 \text{ atm} \times 0.25) = 1.35 \text{ atm}$$

$$V_2 = \frac{V_1 P_1}{P_2} = \frac{(145.7 \text{ mL})(1.08 \text{ atm})}{1.35 \text{ atm}} = 117 \text{ mL}$$

Problem-Solving Lab

page 444



1. **Apply** Boyle's law to explain why air enters your lungs when you inhale and leaves when you exhale.

Boyle's Law states that at constant temperature, the volume of a gas varies inversely with pressure. When you inhale, lung volume increases. Pressure decreases and air moves in. When you exhale, lung volume decreases. Pressure increases and air moves out.

2. **Explain** what happens inside the lungs when a blow to the abdomen knocks the wind out of a person. Use Boyle's law to determine your answer.

When someone is hit in the abdomen, the diaphragm is temporarily paralyzed. When the diaphragm does not move up and down, lung volume does not change. If the volume does not change, pressure inside the lungs does not change, and air is not drawn in and out of the lungs.

3. **Infer** Parts of the lungs lose elasticity and become enlarged when a person has emphysema. From what you know about Boyle's law, why does this condition affect breathing?

If parts of the lungs lose elasticity, it is more difficult to change lung volume. With less change in volume, there is less difference in pressure. With less difference in pressure, it is more difficult for air to enter or exit the lungs.

Gases Solutions Manual Chemistry

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