

Honors Chemistry Rate Law Homework

- 1) For $2A + B \rightarrow C$, we've determined the following experimental data:

Trial	[A] (M)	[B] (M)	Initial rate (M/s)
1	0.0100	0.0100	1.62×10^{-5}
2	0.0200	0.0100	3.24×10^{-5}
3	0.0100	0.0200	6.48×10^{-5}

Using this information, determine the rate law for this reaction, find the reaction order for A and B, and find the overall reaction order.

- 2) In the reaction $A + B \rightarrow C$, we found the following rate data:

Trial	[A] (M)	[B] (M)	Initial rate (M/s)
1	0.026	0.015	2.80×10^{-3}
2	0.026	0.030	5.60×10^{-3}
3	0.052	0.015	11.2×10^{-3}

Using this information, find the overall rate law for this reaction, find the order of the reaction for each reactant, and the overall reaction order.

- 3) Using the information from the rate equation you found in problem 2, determine the rate constant k for this reaction.

Honors Chemistry Rate Law Homework

RS Peters



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