

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify a conjugate acid-base pair in chemical reactions involving acids and bases.
2. Calculate the pH of a solution from the concentration of H^+ or H_3O^+ ions, and vice versa.
3. Determine the pH of a solution from K_a and K_b .

Vocabulary

acid	1. conjugate base	acid solution
acid-base indicator	2. hydronium	acidic solution
acidic solution	3. autoionization	pH

Key Equation

- a. $pH = -\log [H^+]$ (ppm 1.4.1 and 1.4.2)
- b. $[H^+] = 10^{-pH}$ (ppm 1.4.1 and 1.4.2)

Part A Completion

Write the answers to the following questions in the spaces provided. Show necessary calculations and units. Indicate whether the answer is greater or less than, smaller or larger, or halfway.

1. Acetic acid has a K_a of 1.8×10^{-5} . Is it a strong or weak acid?
2. The pH of a solution is 4.5. What is the concentration of H^+ ions?
3. The pH of a solution is 10.5. What is the concentration of OH^- ions?
4. The pH of a solution is 4.5. What is the concentration of OH^- ions?
5. The pH of a solution is 4.5. What is the concentration of H^+ ions?
6. The pH of a solution is 4.5. What is the concentration of OH^- ions?
7. The pH of a solution is 4.5. What is the concentration of H^+ ions?
8. The pH of a solution is 4.5. What is the concentration of OH^- ions?
9. The pH of a solution is 4.5. What is the concentration of H^+ ions?
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14. The pH of a solution is 4.5. What is the concentration of OH^- ions?
15. The pH of a solution is 4.5. What is the concentration of H^+ ions?
16. The pH of a solution is 4.5. What is the concentration of OH^- ions?
17. The pH of a solution is 4.5. What is the concentration of H^+ ions?
18. The pH of a solution is 4.5. What is the concentration of OH^- ions?
19. The pH of a solution is 4.5. What is the concentration of H^+ ions?
20. The pH of a solution is 4.5. What is the concentration of OH^- ions?

Hydrogen Ions And Acidity Section Review Answers

M Tight



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Table of Contents Hydrogen Ions And Acidity Section Review Answers

1. Understanding the eBook Hydrogen Ions And Acidity Section Review Answers
 - The Rise of Digital Reading Hydrogen Ions And Acidity Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hydrogen Ions And Acidity Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hydrogen Ions And Acidity Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hydrogen Ions And Acidity Section Review Answers
 - Personalized Recommendations
 - Hydrogen Ions And Acidity Section Review Answers User Reviews and Ratings
 - Hydrogen Ions And Acidity Section Review Answers and Bestseller Lists
5. Accessing Hydrogen Ions And Acidity Section Review Answers Free and Paid eBooks
 - Hydrogen Ions And Acidity Section Review Answers Public Domain eBooks
 - Hydrogen Ions And Acidity Section Review Answers eBook Subscription Services
 - Hydrogen Ions And Acidity Section Review Answers Budget-Friendly Options

6. Navigating Hydrogen Ions And Acidity Section Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Hydrogen Ions And Acidity Section Review Answers Compatibility with Devices
 - Hydrogen Ions And Acidity Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hydrogen Ions And Acidity Section Review Answers
 - Highlighting and Note-Taking Hydrogen Ions And Acidity Section Review Answers
 - Interactive Elements Hydrogen Ions And Acidity Section Review Answers
8. Staying Engaged with Hydrogen Ions And Acidity Section Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hydrogen Ions And Acidity Section Review Answers
9. Balancing eBooks and Physical Books Hydrogen Ions And Acidity Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hydrogen Ions And Acidity Section Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Hydrogen Ions And Acidity Section Review Answers
 - Setting Reading Goals Hydrogen Ions And Acidity Section Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Hydrogen Ions And Acidity Section Review Answers
 - Fact-Checking eBook Content of Hydrogen Ions And Acidity Section Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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