

Practice Test Ch 3

Stoichiometry

Name Key No.

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" - common factors and rough estimation - do not do "long division" to try to get exact values. Remember it is a MC test, use the answers.
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of NH_3 requires mole(s) of O_2 .

- a. 0.57
b. 1.25
c. 1.33
d. 1.75
e. 3.5

$$1 \text{ mol NH}_3 \left(\frac{7 \text{ mol O}_2}{4 \text{ mol NH}_3} \right) =$$

2. Write a balanced equation for the combustion of acetaldehyde,
- CH_3CHO
- .



When properly balanced, the equation indicates that mole(s) of O_2 are required for each mole of CH_3CHO .

- a. 1
b. 2
c. 2.5
d. 3
e. 6

$$1 \text{ mol CH}_3\text{CHO} \left(\frac{5 \text{ mol O}_2}{2 \text{ mol CH}_3\text{CHO}} \right) =$$

3. What is the total mass of products formed when 16 grams of
- CH_4
- is burned with excess oxygen?



$$16 \text{ g CH}_4 \left(\frac{1 \text{ mol CH}_4}{16 \text{ g CH}_4} \right) \left(\frac{2 \text{ mol O}_2}{1 \text{ mol CH}_4} \right) \left(\frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) =$$

- a. 32 g
b. 36 g
c. 44 g
d. 62 g
e. 80 g

4. Write a balanced equation for the combustion of propane,
- C_3H_8
- .



When balanced, the equation indicates that moles of O_2 are required for each mole of C_3H_8 .

- a. 1.5
b. 3
c. 2.5
d. 5
e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the sum of the coefficients in the balanced equation:



- a. 3
b. 5
c. 6
d. 7
e. 8

6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below:



- a. 1.5 g
b. 2.0 g
c. 3.0 g
d. 6.0 g
e. 12 g

$$27 \text{ g Al} \left(\frac{1 \text{ mol Al}}{27 \text{ g Al}} \right) \left(\frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \right) \left(\frac{2 \text{ g H}_2}{1 \text{ mol H}_2} \right) =$$

7. How many grams of nitric acid,
- HNO_3
- , can be prepared from the reaction of 138 g of
- NO_2
- with 54.0 g
- H_2O
- according to the equation below?



$$138 \text{ g NO}_2 \left(\frac{1 \text{ mol NO}_2}{46 \text{ g NO}_2} \right) \left(\frac{2 \text{ mol HNO}_3}{3 \text{ mol NO}_2} \right) \left(\frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right) = 126 \text{ g}$$

- a. 92
b. 108
c. 126
d. 189
e. 279

8. Which of the following statements is true?

- I. The molar mass of CaCO_3 is 100.1 g mol^{-1} .
II. 50 g of CaCO_3 contains 9×10^{23} oxygen atoms.
III. A 200 g sample of CaCO_3 contains 2 moles of CaCO_3 .

- a. I only
b. II only
c. III only
d. I and III only
e. I, II, and III

$$\frac{40}{100.1} \text{ mol}$$

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Martha B. Barrett



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