

- A. $\text{H}_2\text{O}(g) + 2\text{CO}(g) \rightarrow \text{heat} + 2\text{H}_2\text{O}(l)(g)$
- B. $\text{H}_2\text{O}(g) + 2\text{H}_2\text{O}(g) \rightarrow 2\text{H}_2\text{O}(g) + \text{heat}$
- C. $\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightarrow \text{heat} + \text{CO}(g) + 2\text{H}_2\text{O}(g)$
- D. $\text{C}_2\text{H}_2(g) + 4\text{H}_2(g) \rightarrow \text{CH}_4(g) + 2\text{H}_2\text{S}(g) + \text{heat}$

26. Consider the following equilibrium system: $\text{H}_2(g) + 2\text{H}_2\text{O}(g) \rightleftharpoons 2\text{H}_2\text{O}(g) + 92 \text{ kJ}$. In order to maximize the yield for this reaction, the best conditions are:

- A. Low pressure and low temperature
- B. Low pressure and high temperature
- C. High temperature and low pressure
- D. High pressure and low temperature

27. Consider the following equilibrium: $\text{CaCO}_3(s) \rightleftharpoons 556 \text{ kJ} \rightleftharpoons \text{CaO}(s) + \text{CO}_2(g)$. The equilibrium will shift to the right:

- A. CO_2 is added
- B. $\text{CaCO}_3(s)$ is added
- C. The temperature is increased
- D. The temperature is decreased.

28. Explain the difference between Arrhenius acids and bases and Brønsted-Lowry acids and bases.

29. List three characteristics that acids and bases have in common.

30. To raise the pH of a solution, would you add HCl or NaOH? Why?

31. Describe and give an example of an electrolyte.

32. Find the percent composition of each element in $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$.

33. Given the balanced equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of H_2O can be produced from 50g of O_2 ?

34. Draw energy diagrams for the following:
 Reactants 50kJ / Products 20kJ
 Products 30kJ / Products 40kJ
 Activation energy 20kJ / Activation energy 50kJ

35. Given the unbalanced equation below, what mass of sodium chloride can be produced from 23 g of Na and 35g of Cl_2 ?
 $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$

36. Explain the difference between saturated and unsaturated carbon compounds.

37. Identify the nuclear forces and explain each.

38. Explain how chemical reactions differ from nuclear reactions.

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