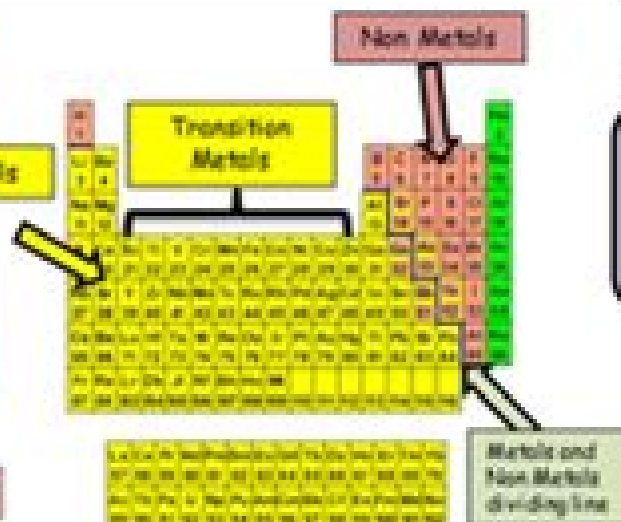


Year 11 C2 - Discovering Chemistry Topic 1

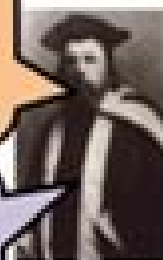


Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-



I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

Dmitri Mendeleev



I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

Isotopes

Higher

Example is an atom with a different number of neutrons

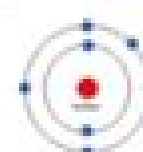
Notice that the mass number is different. How many neutrons does each isotope have?

16 17 18

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more

N

14
7



Electron structure
+ 2, 5

Mg

24
12



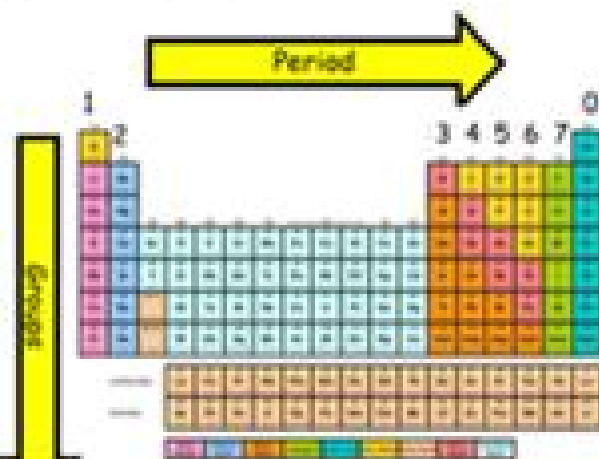
Electron structure
+ 2, 8, 2

Ca

40
20



Electron structure
+ 2, 8, 8, 2

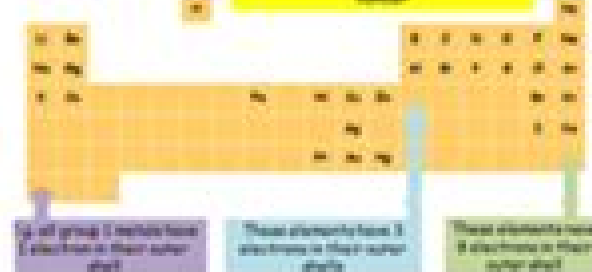


How can an atom have a decimal for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$$35 \times 3 + 1 \times 37 = 106$$

Elements in the same group have the same number of electrons in the outer shell (this corresponds to their group number)



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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Léo Pomar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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