

PART A –CHEMISTRY

1. An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism:
 (1) tertiary alcohol by S_N1 (2) secondary alcohol by S_N2
 (3) tertiary alcohol by S_N2 (4) secondary alcohol by S_N1

Sol. (1)
 Reaction proceeds through carbocation formation as 3° carbocation is highly stable, hence reaction proceeds through S_N1 with 3° alcohol.

2. The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of Na^+ will be:
 (1) - 5.1 eV (3) -10.2 eV
 (2) + 2.55 eV (4) - 2.55 eV

Sol. (1)
 $Na \xrightleftharpoons[\Delta H = -5.1 \text{ eV}]{\Delta H = +5.1 \text{ eV}} Na^+ + e^-$, here the backward reaction releases same amount of energy and known as Electron gain enthalpy.

3. Stability of the species Li_2 , Li_2^+ and Li_2^- increases in the order of:
 (1) $Li_2^- < Li_2^+ < Li_2$ (3) $Li_2 < Li_2^- < Li_2^+$
 (2) $Li_2^+ < Li_2 < Li_2^-$ (4) $Li_2 < Li_2^+ < Li_2^-$

Sol. (1)
 $Li_2 (6) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^2$
 $B.O. = \frac{4-2}{2} = 1$
 $Li_2^+ (5) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^1$
 $B.O. = \frac{3-2}{2} = 0.5$
 $Li_2^- (7) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^2 \sigma^* 2s^1$
 $B.O. = \frac{4-3}{2} = 0.5$
 Li_2^+ is more stable than Li_2^- because Li_2^+ has more numbers of antibonding electrons.

4. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
 (1) 1.00 M (2) 1.75 M
 (3) 0.975 M (4) 0.875 M

Sol. (4)
 $M_1V_1 + M_2V_2 = MV$
 $M = \frac{M_1V_1 + M_2V_2}{V} = \frac{0.5 \times 750 + 2 \times 250}{1000}$
 $M = 0.875$

5. Which of the following is the wrong statement?
 (1) O_3 molecule is bent (2) Ozone is violet-black in solid state
 (3) Ozone is diamagnetic gas (4) $ONCl$ and ONO^+ are not isoelectronic

Sol. (All the options are correct statements)

(1) Correct, as  O_3 is bent.

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Roman Wölfel



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Table of Contents Jee Main Paper Answers 2013

1. Understanding the eBook Jee Main Paper Answers 2013
 - The Rise of Digital Reading Jee Main Paper Answers 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Jee Main Paper Answers 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jee Main Paper Answers 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Jee Main Paper Answers 2013
 - Personalized Recommendations
 - Jee Main Paper Answers 2013 User Reviews and Ratings
 - Jee Main Paper Answers 2013 and Bestseller Lists

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 - Jee Main Paper Answers 2013 Public Domain eBooks
 - Jee Main Paper Answers 2013 eBook Subscription Services
 - Jee Main Paper Answers 2013 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Jee Main Paper Answers 2013 Compatibility with Devices
 - Jee Main Paper Answers 2013 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jee Main Paper Answers 2013
 - Highlighting and Note-Taking Jee Main Paper Answers 2013
 - Interactive Elements Jee Main Paper Answers 2013
8. Staying Engaged with Jee Main Paper Answers 2013
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jee Main Paper Answers 2013
9. Balancing eBooks and Physical Books Jee Main Paper Answers 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jee Main Paper Answers 2013
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Jee Main Paper Answers 2013
 - Setting Reading Goals Jee Main Paper Answers 2013
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Jee Main Paper Answers 2013
 - Fact-Checking eBook Content of Jee Main Paper Answers 2013
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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