

5. For the equilibrium, $A(g) \rightleftharpoons B(g)$, ΔH is -40 kJ/mol . If the ratio of the activation energies of the forward (E_f) and reverse

(E_b) reactions is $\frac{2}{3}$ then :

- (1) $E_f = 60 \text{ kJ/mol}$; $E_b = 100 \text{ kJ/mol}$
- (2) $E_f = 30 \text{ kJ/mol}$; $E_b = 70 \text{ kJ/mol}$
- (3) $E_f = 80 \text{ kJ/mol}$; $E_b = 120 \text{ kJ/mol}$
- (4) $E_f = 70 \text{ kJ/mol}$; $E_b = 30 \text{ kJ/mol}$

6. Determination of the molar mass of acetic acid in benzene using freezing point depression is affected by :

- (1) dissociation
- (2) association
- (3) partial ionization
- (4) complex formation

7. The increase of pressure on ice = water system at constant temperature will lead to :

- (1) no effect on the equilibrium
- (2) a decrease in the entropy of the system
- (3) a shift of the equilibrium in the forward direction
- (4) an increase in the Gibbs energy of the system

8. At 298 K , the standard reduction potentials are 1.51 V for $\text{MnO}_4^-/\text{Mn}^{2+}$, 1.36 V for Cl_2/Cl^- , 1.07 V for Br_2/Br^- , and 0.54 V for I_2/I^- . At $\text{pH} = 3$, permanganate is expected to oxidize :

$$\frac{RT}{F} = 0.059 \text{ V}$$

- (1) Cl^- , Br^- and I^-
- (2) Cl^- and Br^-
- (3) Br^- and I^-
- (4) I^- only

5. इस साम्यावस्था के लिए, $A(g) \rightleftharpoons B(g)$, ΔH -40 kJ/mol है। अगर अग्र (E_f) और पश्चगामी (E_b) अभिक्रियाओं के सक्रियण ऊर्जाओं का अनुपात

$\frac{2}{3}$ है, तब :

- (1) $E_f = 60 \text{ kJ/mol}$; $E_b = 100 \text{ kJ/mol}$
- (2) $E_f = 30 \text{ kJ/mol}$; $E_b = 70 \text{ kJ/mol}$
- (3) $E_f = 80 \text{ kJ/mol}$; $E_b = 120 \text{ kJ/mol}$
- (4) $E_f = 70 \text{ kJ/mol}$; $E_b = 30 \text{ kJ/mol}$

6. ऐसिटिक अम्ल का मोलर द्रव्यमान बेंजीन में हिमांक अवनमन से निकाले जाने पर जिससे प्रभावित होता है वह है :

- (1) वियोजन
- (2) संगुणन
- (3) आंशिक आयनीकरण
- (4) संकुल बिरचन

7. स्थिर ताप पर, अधिक दाब डालने पर बर्फ = जल समुदाय में :

- (1) साम्यावस्था पर कोई प्रभाव नहीं होता है।
- (2) निकाय की ऐन्ट्रॉपी घटती है।
- (3) साम्यावस्था अग्र दिशा में अग्रसर होती है।
- (4) निकाय की गिब्स ऊर्जा बढ़ती है।

8. 298 K पर मानक अपचयन विभव, $\text{MnO}_4^-/\text{Mn}^{2+}$ का 1.51 V , Cl_2/Cl^- का 1.36 V , Br_2/Br^- का 1.07 V और I_2/I^- का 0.54 V है। $\text{pH} = 3$ पर MnO_4^- किसका अपचयन करेगा ?

$$\frac{RT}{F} = 0.059 \text{ V}$$

- (1) Cl^- , Br^- और I^-
- (2) Cl^- और Br^-
- (3) Br^- और I^-
- (4) I^- मात्र

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