

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্লো, এখানে মোট পটাসিয়াম, আরও কপ্লোয়াম পটাসিয়ামের পটাসিয়াম এবং মনোমের পটাসিয়াম
পটাসিয়াম উত্তর অধ্যায়িকিত সেএক্স হলেও। এপুলের অনুবীলন করলে তুমি এ অধ্যায়িকিত থেকে সেএক্সের পটাসিয়াম পটাসিয়াম
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পটাসিয়াম লবণের প্রকণতা, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রকণ হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেন্দ্র পটাসিয়াম লবণের প্রকণতা অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেন্দ্র পটাসিয়াম লবণ নিষ্কৃতি হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পটাসিয়ামে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি 32.4°C আপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার হ্রাস ঘটে।

প্রশ্ন ২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন ১: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন ২: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

প্রশ্ন ৩: A-পাত্রের প্রকণের সাথে প্রশ্ন ২-এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ৪: একমাত্র প্রকণের উপরে X-পটাসিয়ামে 10 mg এর প্রকণ প্রকণ এবং Y-পটাসিয়ামে 0.5 g পরিমাণ 2% $\frac{w}{v}$ HCl এর প্রকণ প্রকণ করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোৎপাদী— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পটাসিয়ামে পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

উদাহরণের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

উদাহরণের সমযোজী বন্ধনে আবদ্ধ উদাহরণ পরমাণুর তড়িৎ ঋণাত্মকতা হল অর্ধাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বৃদ্ধি ও অপোৎপাদী হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোৎপাদী হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপোৎপাদী তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বৃদ্ধি সমযোজী ও অপোৎপাদী যৌগ। উদাহরণ মৌলিক অণুর ক্ষেত্রে এহুণ থাকে।

প্রশ্ন ১: উদাহরণের বর্ণিত এলিটটি হলো 2% $\left(\frac{w}{v}\right)$ HCl।

2% $\left(\frac{w}{v}\right)$ HCl এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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CL Gary



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