16 • Chemical Equilibrium

Consider the reaction system,

 $CoO(s) + H_2(g) \stackrel{\text{def}}{\leftarrow} Co(s) + H_2O(g)$.

The equilibrium constant expression is

- [CoO][H.] [Co][H.O]
- [H.O]
- $[Co][H_2O]$ [CoO][H.] Col[H.O] HI.
- Given the equilibrium.

2SO₂(g) + O₂(g) ∓ 2SO₂(g), if this equilibrium

is established by beginning with equal number of moles of SO₂ and O₂ in a 1.0 Liter bulb, then

the following must be true at equilibrium:

- a) [SO₂] = [SO₃]d) [SO₂] < [O₂]
- (b) $2|SO_2| = 2|SO_3|$ e) $|SO_2| > |O_2|$
 - c) [SO₂] = [O₂]

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO, 0.200 mol Cb and 0.500 mol CINO were placed in a 25.0 Liter container. The following equilibrium is established: 2CINO(g) ₹ 2NO(g) + Cb(g)

- At equilibrium, 0.600 mol of CINO was present. The number of moles of Cb present at equilibrium is
 - cal 0.050

e) 0.250

b) 0.100 c) 0.150

- PRACTICE TEST
- The equilibrium constant, Kc, is:
 - a) 4.45 x 10⁻⁴
- d) 0.167
- b) 6.67 x 10⁻⁴
- e) 1500

- ch 0.111
- At 985°C, the equilibrium constant for the reaction.

$$H_2(g) + CO_2(g) \rightleftharpoons H_2O(g) + CO(g)$$

is 1.63. What is the equilibrium constant for the reverse reaction?

a) 1.63

d) 0.613

b) 0.815

e)) 1.00

- c) 2.66
- What is the relationship between K_p and K_c for the reaction, $2ICI(g) \rightleftharpoons I_2(g) + Cb(g)$?
 - (a) $k_p = K_c(RT)^{-1}$ (d) $K_p = K_c$

- b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
- c) $K_p = K_c(RT)^2$
- For the reaction 2NO₂(g) ₹ N₂O₄(g), K_n at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is Ke for this reaction? [R=0.0821 Latmmol* K*]

4270

d) 179

b) 0.0119

e) = 2.06

c) = 0.291

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ML Morrison

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Table of Contents Honors Chemistry Equilibrium Practice

- 1. Understanding the eBook Honors Chemistry Equilibrium Practice
 - The Rise of Digital Reading Honors Chemistry Equilibrium Practice
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Honors Chemistry Equilibrium Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Equilibrium Practice
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Honors Chemistry Equilibrium Practice
 - Personalized Recommendations
 - Honors Chemistry Equilibrium Practice User Reviews and Ratings

- Honors Chemistry Equilibrium Practice and Bestseller Lists
- 5. Accessing Honors Chemistry Equilibrium Practice Free and Paid eBooks
 - Honors Chemistry Equilibrium Practice Public Domain eBooks
 - Honors Chemistry Equilibrium Practice eBook Subscription Services
 - Honors Chemistry Equilibrium Practice Budget-Friendly Options
- 6. Navigating Honors Chemistry Equilibrium Practice eBook Formats
 - o ePub, PDF, MOBI, and More
 - Honors Chemistry Equilibrium Practice Compatibility with Devices
 - Honors Chemistry Equilibrium Practice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Honors Chemistry Equilibrium Practice
 - Highlighting and Note-Taking Honors Chemistry Equilibrium Practice
 - Interactive Elements Honors Chemistry Equilibrium Practice
- 8. Staying Engaged with Honors Chemistry Equilibrium Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Equilibrium Practice
- 9. Balancing eBooks and Physical Books Honors Chemistry Equilibrium Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Equilibrium Practice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Honors Chemistry Equilibrium Practice
 - Setting Reading Goals Honors Chemistry Equilibrium Practice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Honors Chemistry Equilibrium Practice
 - Fact-Checking eBook Content of Honors Chemistry Equilibrium Practice
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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