

Math Physics

Problem 2B

PROBLEM 2B (20 points)

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The power spectrum calculated is shown in the graph below. What wavelength (nm) of the incident light is responsible for the observed electron emission? (100% electron kinetic energy conversion speed efficiency assumed)

PROBLEM 2B (20 points)

Given:

100% electron kinetic energy conversion efficiency
 $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$

Calculate:

20% (10)

Energy spectrum of incident electromagnetic wave:

$W_{\text{max}} = 10 \text{ eV}$
 $W_{\text{min}} = 0 \text{ eV}$

Wavelength spectrum of incident wave:

$\lambda_{\text{max}} = 1240 \text{ nm}$
 $\lambda_{\text{min}} = 1240 \text{ nm}$
 $\lambda_{\text{max}} = 1240 \text{ nm}$
 $\lambda_{\text{min}} = 1240 \text{ nm}$

100%

ANSWER KEY (20 points)

1. The incident light wavelength is 1240 nm, which will be equal to the wavelength of the incident light when the electron kinetic energy is 100%.
2. The incident light wavelength is 1240 nm, which will be equal to the wavelength of the incident light when the electron kinetic energy is 100%.
3. The incident light wavelength is 1240 nm, which will be equal to the wavelength of the incident light when the electron kinetic energy is 100%.

Holt Physics Practice 15a

VM Jensen

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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