

What's New? A new simulation about Molecular Geometry and Polarity.

A study of electronegativity, bond polarity, and molecular polarity (part 1)

Introduction

In this atomic-level simulation, you will investigate how atoms' electronegativity value affects the bond's behavior. When two atoms bond, a pair of electrons is shared between atoms. Electronegativity is a measure of a single atom's ability to hold electrons shared in that bond. In this lab you will work diligently, at your own pace, to answer a number of questions.

If a ligand, from what you've already learned about the protons and electrons in an atom, what kind of gases are atoms to have a high electronegativity values? high number of valence electrons, mean an atom also small atoms with low shielding.

Why might an atom have a low electronegativity value?

atoms with more than one shell (3, 4, 5) large atoms with lots of shielding.

Procedures

- Turn on (click) all view options: ORBIT and spin the molecule with your mouse.
- Play your time and investigate how the binary compound's bond behaves when the atom's electronegativity and oxidation are changed. Do not rush through this step.

Describe the polarity of the bond formed between two atoms with similar, low electronegativity.

____nonpolar, no molecular dipole____

Describe the bond formed between two atoms with similar, high electronegativity.

____nonpolar, no molecular dipole____

Describe the bond formed between two atoms with very different electronegativity.

____polar, has molecular dipole____

Describe (in your own words) what is meant by partial charges, δ^- and δ^+ . ____positive has or negative has____

δ^- represent to ____partial negative____

δ^+ represent to ____partial positive____

What happens when the electric field is applied to a very polar molecule? ____molecule align itself with the field____

Why do you think this is? ____opposites attract and likes repel____

What is electron density? ____electron cloud is more attracted to one atom or side of a molecule____

How does the density around a partial positive compare to the density of a partial negative? ____

____partial positive density < partial negative density____

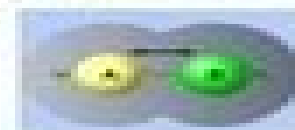
What would bring about a higher electron density around an atom? ____higher electronegativity value____

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