Key

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 - Lewis Structures

Lows Structures	H ₂ CO 3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
2. Ref. #06#	BeF ₂ 2 electron domains (2 bonding, 0 nonbonding)	2 Tetrakedra
3.04	CH, 4 electron domains (4 bonding, 0 nonbonding)	o go Tetrakedra
5. H ₂ O	NH ₃ 4 electron domains (3 bonding, 1 nonbonding)	800 pyrami
·.o,	H ₃ O 4 electron domains (2 bonding, 2 nonbonding)	a Bent
Lone pair - ·-	CO ₂ 2 electron domains (2 bonding, 0 nonbonding)	O=O=O

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