

Key

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are *three-dimensional*. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures		3-D Molecular Shape
1. H_2CO	3 electron domains (3 bonding, 0 nonbonding)	trigonal planar
2. BeF_2	2 electron domains (2 bonding, 0 nonbonding)	Linear
3. CH_4	4 electron domains (4 bonding, 0 nonbonding)	Tetrahedral
4. NH_3	4 electron domains (3 bonding, 1 nonbonding)	pyramidal
5. H_2O	4 electron domains (2 bonding, 2 nonbonding)	Bent
6. CO_2	2 electron domains (2 bonding, 0 nonbonding)	linear

Lone pair = ..

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