

## Empirical and Molecular Formula Practice KEY

### Empirical Formula

1. 36.5% Na, 25.4% S, and 38.1% O forms this compound.

$$\text{Na} - 36.5/23 = 1.56/0.8 = 2$$

$$\text{S} - 25.4/32 = 0.8/0.8 = 1$$

$$\text{O} - 38.1/16 = 2.4/0.8 = 3$$



2. A compound made of 72.2% Magnesium and 27.8% nitrogen by mass.

$$\text{Mg} - 72.2/24.3 = 3/2 = 1.5(2) = 3$$

$$\text{N} - 27.8/14 = 2/2 = 1(2) = 2$$



3. A chromium compound with 68.4% Cr and 31.6% O.

$$\text{Cr} - 68.4/52 = 1.315/1.315 = 1(2) = 2$$

$$\text{O} - 31.6/16 = 1.975/1.315 = 1.5(2) = 3$$



4. A compound with the composition of 24% C, 7% H, 38% F, and 31% P.

$$\text{C} - 24/12 = 2/1 = 2$$

$$\text{H} - 7/1 = 7/1 = 7$$

$$\text{F} - 38/19 = 2/1 = 2$$

$$\text{P} - 31/31 = 1/1 = 1$$



5. 63.5% silver, 8.2% nitrogen, and 28.3% oxygen make up this compound.

$$\text{Ag} - 63.5/107.87 = 0.589/0.585 = 1$$

$$\text{N} - 8.2/14.01 = 0.585/0.585 = 1$$

$$\text{O} - 28.3/16 = 1.769/0.585 = 3$$



6. 8.81g C and 91.2g Chlorine forms this compound.

$$\text{C} - 8.81/12 = 0.73/0.73 = 1(2) = 2$$

$$\text{Cl} - 91.2/35 = 2.61/0.73 = 3.5(2) = 7$$



7. A compound in a soft drink that has 0.31g hydrogen, 3.2g phosphorus, and 6.5g oxygen.

$$\text{H} - 0.31/1 = 0.31/.1 = 3.1$$

$$\text{P} - 3.2/31 = 0.10/.1 = 1$$

$$\text{O} - 6.5/16 = 0.40/.1 = 4$$



# Molecular And Empirical Formula Practice

**R Sandford**



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## **Table of Contents Molecular And Empirical Formula Practice**

1. Understanding the eBook Molecular And Empirical Formula Practice
  - The Rise of Digital Reading Molecular And Empirical Formula Practice
  - Advantages of eBooks Over Traditional Books
2. Identifying Molecular And Empirical Formula Practice
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molecular And Empirical Formula Practice
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular And Empirical Formula Practice
  - Personalized Recommendations
  - Molecular And Empirical Formula Practice User Reviews and Ratings
  - Molecular And Empirical Formula Practice and Bestseller Lists
5. Accessing Molecular And Empirical Formula Practice Free and Paid eBooks

- Molecular And Empirical Formula Practice Public Domain eBooks
  - Molecular And Empirical Formula Practice eBook Subscription Services
  - Molecular And Empirical Formula Practice Budget-Friendly Options
6. Navigating Molecular And Empirical Formula Practice eBook Formats
    - ePub, PDF, MOBI, and More
    - Molecular And Empirical Formula Practice Compatibility with Devices
    - Molecular And Empirical Formula Practice Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Molecular And Empirical Formula Practice
    - Highlighting and Note-Taking Molecular And Empirical Formula Practice
    - Interactive Elements Molecular And Empirical Formula Practice
  8. Staying Engaged with Molecular And Empirical Formula Practice
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Molecular And Empirical Formula Practice
  9. Balancing eBooks and Physical Books Molecular And Empirical Formula Practice
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Molecular And Empirical Formula Practice
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Molecular And Empirical Formula Practice
    - Setting Reading Goals Molecular And Empirical Formula Practice
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Molecular And Empirical Formula Practice
    - Fact-Checking eBook Content of Molecular And Empirical Formula Practice
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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