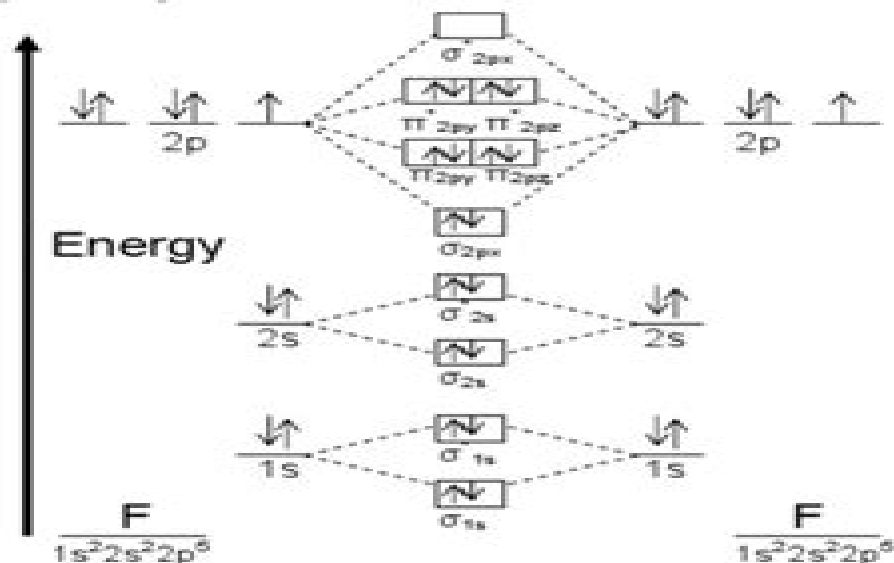


These problems are for practice in drawing your molecular orbital diagrams, molecular electron configurations and determining bond order.

The following questions pertain to the F_2 molecule:

- A) Draw the molecular orbital energy diagram for this molecule. Label all of the orbitals specifically.



- B) Give the molecular electron configuration for the molecule:



- C) Determine the bond order for the molecule:

$$\text{Bond Order} = (10 - 8) / 2 = 1$$

- D) Indicate whether the species is paramagnetic or diamagnetic:

Diamagnetic: No unpaired electrons

- F) Compare the relative stability of this molecule to F_2^+ and F_2^- :



Molecular Orbital Practice Answers

R Barnett



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